

## Covalent and Metallic Bonds

### BEFORE YOU READ

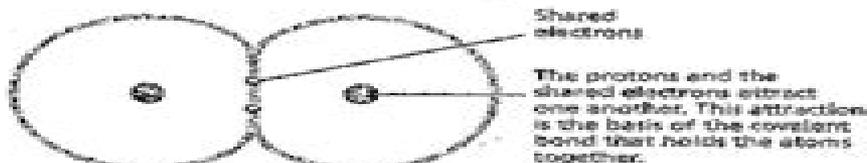
After you read this section, you should be able to answer these questions:

- What are covalent bonds?
- What are molecules?
- What are metallic bonds?
- How does bonding affect a metal's properties?

### What Are Covalent Bonds?

Another type of bond is a covalent bond. A **covalent bond** forms when atoms share electrons. Covalent bonds most often form between atoms of nonmetals. Remember that most nonmetals can fill the outermost energy level by gaining an electron. When a covalent bond forms, both atoms are able to fill their outermost energy level. They do this by sharing electrons between the two atoms.  $\text{R}$

Hydrogen is one example of an atom that bonds covalently. A hydrogen atom has one electron in its outermost level. Two hydrogen atoms can come together and share their electrons. This fills the first energy level of both atoms. The electrons move around both hydrogen nuclei. The protons and the shared electrons attract one another. This attraction holds the atoms together.



By sharing electrons in a covalent bond, each hydrogen atom (the smallest atom) has a full outermost energy level containing two electrons.

### What Are Molecules?

Atoms that join with each other by covalent bonds form particles called **molecules**. Most molecules are made of atoms of two or more elements. The atoms share electrons. In the figure above, two hydrogen atoms have formed a covalent bond. The result is a hydrogen molecule.  $\text{R}$

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### STUDY TIP

**Compare** As you read, make a chart comparing covalent bonds and metallic bonds.

### READING CHECK

1. Explain How do electrons behave in covalent bonds?

The electrons are shared.

### READING CHECK

2. Identify What type of bond joins the atoms in molecules?

Covalent Bonds

# Guided Study Covalent Bonds Answer Key

**Katherine J. Denniston, Denniston  
Katherine**



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