

Mole calculations – guided practice

Question 1



What mass of ammonia (NH_3) will be made if I react 5.6 g of nitrogen (N_2) with an excess of hydrogen gas?

Step 1:

Work out the **formula mass** of N_2 by finding the atomic mass of nitrogen (N) on the periodic table and multiplying it by 2.

Formula mass of N_2

Step 2:

Work out the number of **moles** of N_2 in 5.6 g by using the equation:

$$\begin{aligned} \text{moles} &= \text{mass} / \text{formula mass} \\ &= 5.6 \div \text{your answer to step 1} \end{aligned}$$

Moles of N_2

Step 3: *using the mole ratio in the balanced equation*

If you look at the equation, you should see that **1** mole of N_2 produces **2** moles of NH_3

Guided Practice Problems 1stoichiometry

Peipei Pang



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