

Chemical Bonding (Chapter 8) Section 2

8. How do hybrid orbitals differ from s orbitals?

Hybrid orbitals contain the same amount of s and p character as the s and p orbitals that combine to form them.

9. Is sp^3 hybridization found in molecules that contain double bonds? Explain.

10. Using the bond energies in Table 8.4, what is the enthalpy of formation of ethane, C_2H_6 , from its elements in their standard states?

The enthalpy of formation of ethane from its elements is -84.7 kJ/mol.

11. Which compound has a greater bond energy, C_2H_4 or C_2H_2 ?

12. Are sp^3 hybrid orbitals found in molecules that contain triple bonds?

13. What is a hybrid orbital?

A hybrid orbital is formed by the combination of s and p atomic orbitals.

14. Why are the hybrid orbitals in ethane oriented toward each other?

Hybrid orbitals are oriented toward each other to minimize repulsion.

15. How many electrons does the sp^3 hybrid orbital have the highest energy?

It has one electron.

16. What are the hybrid orbitals in ethane oriented toward each other?

They are oriented toward each other to minimize repulsion.

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