

## Equilibrium Practice Problems

1. Write the equilibrium expression for each of the following reactions:



$$K = \frac{[\text{NH}_3]^2}{[\text{N}_2] [\text{H}_2]^3}$$



$$K = \frac{[\text{ICl}]^2}{[\text{Cl}_2]}$$



$$K = \frac{[\text{NO}] [\text{O}_2]^{1/2}}{[\text{NO}_2]}$$

2. The dissociation of acetic acid,  $\text{CH}_3\text{COOH}$ , has an equilibrium constant at  $25^\circ\text{C}$  of  $1.8 \times 10^{-5}$ . The reaction is



If the equilibrium concentration of  $\text{CH}_3\text{COOH}$  is 0.46 moles in 0.500 L of water and that of  $\text{CH}_3\text{COO}^-$  is  $8.1 \times 10^{-3}$  moles in the same 0.500 L, calculate  $[\text{H}^+]$  for the reaction.

$$K = \frac{[\text{CH}_3\text{COO}^-] [\text{H}^+]}{[\text{CH}_3\text{COOH}]} \quad \text{Plug in known values and solve.}$$

$$1.8 \times 10^{-5} = \frac{[8.1 \times 10^{-3} \text{ moles}/0.500 \text{ L}] [\text{H}^+]}{[0.46 \text{ moles}/0.500 \text{ L}]}$$

$$[\text{H}^+] = 1.0 \times 10^{-3} \text{ M}$$

3. Indicate the effect of a catalyst, pressure, temperature and concentration on each of the following on:

|                     | catalyst          | pressure            | temperature              | concentration        |
|---------------------|-------------------|---------------------|--------------------------|----------------------|
| a. speed            | presence increase | direct relationship | direct relationship      | direct relationship  |
| b. pos <sup>n</sup> | nothing           | based on moles      | based on enthalpy of rxn | inverse relationship |

# Kinetics And Equilibrium Practice Problems

**S Marginson**



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## **Table of Contents Kinetics And Equilibrium Practice Problems**

1. Understanding the eBook Kinetics And Equilibrium Practice Problems
  - The Rise of Digital Reading Kinetics And Equilibrium Practice Problems
  - Advantages of eBooks Over Traditional Books
2. Identifying Kinetics And Equilibrium Practice Problems
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Kinetics And Equilibrium Practice Problems
  - User-Friendly Interface
4. Exploring eBook Recommendations from Kinetics And Equilibrium Practice Problems
  - Personalized Recommendations
  - Kinetics And Equilibrium Practice Problems User Reviews and Ratings
  - Kinetics And Equilibrium Practice Problems and Bestseller Lists
5. Accessing Kinetics And Equilibrium Practice Problems Free and Paid eBooks
  - Kinetics And Equilibrium Practice Problems Public Domain eBooks
  - Kinetics And Equilibrium Practice Problems eBook Subscription Services
  - Kinetics And Equilibrium Practice Problems Budget-Friendly Options
6. Navigating Kinetics And Equilibrium Practice Problems eBook Formats
  - ePub, PDF, MOBI, and More
  - Kinetics And Equilibrium Practice Problems Compatibility with Devices
  - Kinetics And Equilibrium Practice Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Kinetics And Equilibrium Practice Problems
  - Highlighting and Note-Taking Kinetics And Equilibrium Practice Problems
  - Interactive Elements Kinetics And Equilibrium Practice Problems

8. Staying Engaged with Kinetics And Equilibrium Practice Problems
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Kinetics And Equilibrium Practice Problems
9. Balancing eBooks and Physical Books Kinetics And Equilibrium Practice Problems
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Kinetics And Equilibrium Practice Problems
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Kinetics And Equilibrium Practice Problems
  - Setting Reading Goals Kinetics And Equilibrium Practice Problems
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Kinetics And Equilibrium Practice Problems
  - Fact-Checking eBook Content of Kinetics And Equilibrium Practice Problems
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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