NAME: key DATE: _____ PERIOD: _____

GAS STOICHIOMETRY PROBLEMS WORKSHEET 1

1. Given the unbalanced decomposition reaction of baking soda:

$2NaHCO_{3(s)} + \text{heat} \Longrightarrow Na_2CO_{3(s)} + CO_{2(g)} + HOH_{(g)}$

[redacted] 42.0 grams of baking soda? What volumes of carbon dioxide and water are produced at STP?

$42.0\ g\ NaHCO_3 \times \frac{1\ mol\ NaHCO_3}{84.01\ g\ NaHCO_3} \times \frac{1\ mol\ CO_2}{2\ mol\ NaHCO_3} \times \frac{22.4\ L\ CO_2}{1\ mol\ CO_2} = 5.6\ L\ CO_2$

$42.0\ g\ NaHCO_3 \times \frac{1\ mol\ NaHCO_3}{84.01\ g\ NaHCO_3} \times \frac{1\ mol\ CO_2}{2\ mol\ NaHCO_3} \times \frac{22.4\ L\ CO_2}{1\ mol\ CO_2} = 5.6\ L\ CO_2$

2. The catalytic decomposition of hydrogen peroxide is:

$2H_2O_{2(aq)} \Longrightarrow 2H_2O_{(l)} + O_{2(g)}$

Balance the reaction. How many moles of water and oxygen are produced by the decomposition of 68.0 grams of hydrogen peroxide? How many molecules of water and oxygen are produced? How many grams of each product are formed?

$68\ g\ H_2O_2 \times \frac{1\ mol\ H_2O_2}{34.02\ g\ H_2O_2} \times \frac{2\ mol\ H_2O}{2\ mol\ H_2O_2} = 2.00\ mol\ H_2O$

$68\ g\ H_2O_2 \times \frac{1\ mol\ H_2O_2}{34.02\ g\ H_2O_2} \times \frac{1\ mol\ O_2}{2\ mol\ H_2O_2} = 1.00\ mol\ O_2$

$68\ g\ H_2O_2 \times \frac{1\ mol\ H_2O_2}{34.02\ g\ H_2O_2} \times \frac{2\ mol\ H_2O}{2\ mol\ H_2O_2} \times \frac{6.02 \times 10^{23}\ molecules\ H_2O}{1\ mol\ H_2O} = 1.20 \times 10^{24}\ molecules\ H_2O$

$68\ g\ H_2O_2 \times \frac{1\ mol\ H_2O_2}{34.02\ g\ H_2O_2} \times \frac{1\ mol\ O_2}{2\ mol\ H_2O_2} \times \frac{6.02 \times 10^{23}\ molecules}{1\ mol} = 6.02 \times 10^{23}\ molecules\ O_2$

$68\ g\ H_2O_2 \times \frac{1\ mol\ H_2O_2}{34.02\ g\ H_2O_2} \times \frac{2\ mol\ H_2O}{2\ mol\ H_2O_2} \times \frac{18.02\ g\ H_2O}{1\ mol\ H_2O} = 36.02\ g\ H_2O$

$68\ g\ H_2O_2 \times \frac{1\ mol\ H_2O_2}{34.02\ g\ H_2O_2} \times \frac{1\ mol\ O_2}{2\ mol\ H_2O_2} \times \frac{32\ g\ O_2}{1\ mol\ O_2} = 63.94\ g\ O_2$

3. If 156.0 grams of potassium metal reacts with excess water, [redacted] potassium hydroxide and hydrogen gas are formed? What volume of hydrogen gas, in liters, [redacted] Be sure to balance the reaction first.

$2K_{(s)} + 2H_2O_{(l)} \rightarrow 2KOH_{(aq)} + H_{2(g)}$

$156\ g\ K \times \frac{1\ mol\ K}{39.09\ g\ K} \times \frac{1\ mol\ H_2}{2\ mol\ K} \times \frac{22.4\ L\ H_2}{1\ mol\ H_2} = 44.70\ L\ H_2$

$156\ g\ K \times \frac{1\ mol\ K}{39.09\ g\ K} \times \frac{2\ mol\ KOH}{2\ mol\ K} \times \frac{22.4\ L\ KOH}{1\ mol\ KOH} = 89.39\ L\ KOH$

4. Determine the number of moles of carbon dioxide gas, water, and sodium chloride formed by the reaction of 42.0 grams of sodium bicarbonate (baking soda) reacting with excess hydrochloric acid. The reaction is:

$NaHCO_{3(s)} + HCl_{(aq)} \Longrightarrow CO_{2(g)} + HOH_{(l)} + NaCl_{(aq)}$

$42\ g\ NaHCO_3 \times \frac{1\ mol\ NaHCO_3}{84.01\ g\ NaHCO_3} \times \frac{1\ mol\ CO_2}{1\ mol\ NaHCO_3} = 0.5\ mol\ CO_2$

$42\ g\ NaHCO_3 \times \frac{1\ mol\ NaHCO_3}{84.01\ g\ NaHCO_3} \times \frac{1\ mol\ H_2O}{1\ mol\ NaHCO_3} = 0.5\ mol\ H_2O$

$42\ g\ NaHCO_3 \times \frac{1\ mol\ NaHCO_3}{84.01\ g\ NaHCO_3} \times \frac{1\ mol\ NaCl}{1\ mol\ NaHCO_3} = 0.5\ mol\ NaCl$

Gas Stoichiometry Problem And Answer

Lei Shi

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