

AP[®] Chemistry GASES

The gaseous state of matter is the simplest and best-understood state of matter. You inhale approximately 8,500 L of air each day. This amounts to about 25 lbs of air. Breathing is a three-step process: inhaling, gas exchange with the circulatory system, and exhaling. Approximately 80% pressure differences created by your body allow you to breathe. Nearly the Earth's entire atmosphere is made up of only five gases: nitrogen, oxygen, water vapor, argon, and carbon dioxide.

PROPERTIES OF GASES

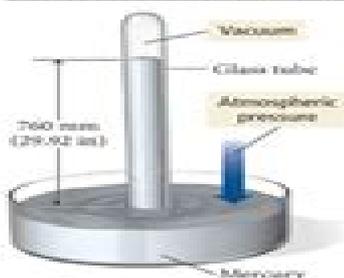
Only **FOUR** quantities are needed to define the state of a gas:

1. the quantity of the gas, n (in moles)
2. the temperature of the gas, T (in **KELVINS**)
3. the volume of the gas, V (in liters)
4. the pressure of the gas, P (in atmospheres)

A gas uniformly fills any container, is easily compressed & mixes completely with any other gas.



The Mercury Barometer



Pressure and Density



GAS PRESSURE is a measure of the force that a gas exerts on its container. Force is the physical quantity that interferes with inertia. Gravity is the force responsible for weight.

Force = mass \times acceleration; (Newton's 2nd Law)

The SI units follow:

$$N = kg \times m/s^2$$

Pressure-- Force/ unit area; N/m^2 ; which is the definition of 1.0 Pascal

Barometer--invented by Evangelista Torricelli in 1643; uses the height of a column of mercury to measure gas pressure (especially atmospheric P)

$$1 \text{ mm of Hg} = 1 \text{ torr}$$

$$1.00 \text{ atm} = 760.00 \text{ mm Hg} = 760.00 \text{ torr} = 29.92 \text{ in Hg} = 14.7 \text{ psi} = 101.325 \text{ kPa} \approx 10^5 \text{ Pa}$$

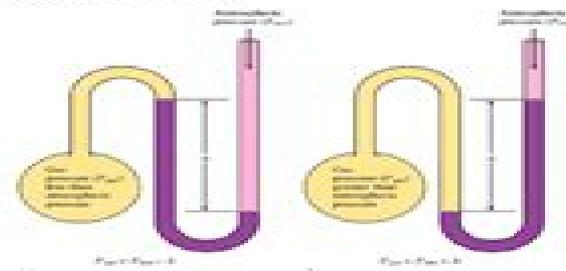
At sea level all of the above define **STANDARD PRESSURE**.

The SI unit of pressure is the Pascal (named after Blaise Pascal);

$$1 \text{ Pa} = 1 \text{ N} / \text{m}^2$$

The **manometer**—a device for measuring the pressure of a gas in a container. The pressure of the gas is given by h [the difference in mercury levels] in units of torr (equivalent to mm Hg).

- a) Gas pressure = atmospheric pressure - h
- b) Gas pressure = atmospheric pressure + h



Gases Packet For Chemistry

Mike Jess



Gases Packet For Chemistry:

Unveiling the Magic of Words: A Report on "**Gases Packet For Chemistry**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Gases Packet For Chemistry**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book's central themes, examine its distinctive writing style, and assess its profound affect on the souls of its readers.

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Table of Contents Gases Packet For Chemistry

1. Understanding the eBook Gases Packet For Chemistry
 - The Rise of Digital Reading Gases Packet For Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Gases Packet For Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gases Packet For Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gases Packet For Chemistry
 - Personalized Recommendations
 - Gases Packet For Chemistry User Reviews and Ratings
 - Gases Packet For Chemistry and Bestseller Lists

5. Accessing Gases Packet For Chemistry Free and Paid eBooks
 - Gases Packet For Chemistry Public Domain eBooks
 - Gases Packet For Chemistry eBook Subscription Services
 - Gases Packet For Chemistry Budget-Friendly Options
6. Navigating Gases Packet For Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Gases Packet For Chemistry Compatibility with Devices
 - Gases Packet For Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gases Packet For Chemistry
 - Highlighting and Note-Taking Gases Packet For Chemistry
 - Interactive Elements Gases Packet For Chemistry
8. Staying Engaged with Gases Packet For Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gases Packet For Chemistry
9. Balancing eBooks and Physical Books Gases Packet For Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gases Packet For Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gases Packet For Chemistry
 - Setting Reading Goals Gases Packet For Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gases Packet For Chemistry
 - Fact-Checking eBook Content of Gases Packet For Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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