

Answers to the Gases Review Sheet

1. 3.21 mm Hg = 0.321 cm Hg = 0.00422 atm
2. 6.9 kPa = 6900 Pa = 52 mm Hg
3. 436 mm Hg = 0.574 atm = 58.1 kPa
4. 2.9 kPa = 22 torr = 22 mm Hg
5. A sample of air has a volume of 2.25 cm³ when its temperature is 298 K. If the temperature is increased to 373 K without changing the pressure, what is the new volume of the sample in air?

$$\begin{aligned} V_1 &= 2.25 \text{ cm}^3 & \frac{V_1}{T_1} &= \frac{V_2}{T_2} & V_2 &= \frac{V_1 T_2}{T_1} \\ T_1 &= 298 \text{ K} \\ T_2 &= 373 \text{ K} \\ V_2 &= x \end{aligned}$$

Charles's

$$V_2 = \frac{2.25 \text{ cm}^3 (373 \text{ K})}{298 \text{ K}} = \mathbf{2.82 \text{ cm}^3}$$

6. As the temperature of a sample of nitrogen increases from 273 K, the volume of the sample changes from 275 mL to 325 mL. Assuming the pressure is constant, what is the final temperature of the gas?

$$\begin{aligned} V_1 &= 275 \text{ mL} & \frac{V_1}{T_1} &= \frac{V_2}{T_2} & T_2 &= \frac{T_1 V_2}{V_1} \\ T_1 &= 273 \text{ K} \\ T_2 &= x \\ V_2 &= 325 \text{ mL} \end{aligned}$$

Charles's

$$T_2 = \frac{325 \text{ mL} (273 \text{ K})}{275 \text{ mL}} = \mathbf{323 \text{ K}}$$

7. The pressure of a 5.71 L sample of neon gas is 23.4 kPa. Calculate the new pressure when the volume becomes 3.40 L.

$$\begin{aligned} V_1 &= 5.71 \text{ L} & P_1 V_1 &= P_2 V_2 & P_2 &= \frac{P_1 V_1}{V_2} \\ P_1 &= 23.4 \text{ kPa} \\ P_2 &= x \end{aligned}$$

$$P_2 = \frac{5.71 \text{ L} (23.4 \text{ kPa})}{3.40 \text{ L}} = 39.3 \text{ kPa} = \mathbf{0.388 \text{ atm}}$$

Boyle's

Gases Section 3 Review Answers

Michael B. Leyden



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Table of Contents Gases Section 3 Review Answers

1. Understanding the eBook Gases Section 3 Review Answers
 - The Rise of Digital Reading Gases Section 3 Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gases Section 3 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gases Section 3 Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gases Section 3 Review Answers
 - Personalized Recommendations
 - Gases Section 3 Review Answers User Reviews and Ratings
 - Gases Section 3 Review Answers and Bestseller Lists
5. Accessing Gases Section 3 Review Answers Free and Paid eBooks

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 - Gases Section 3 Review Answers eBook Subscription Services
 - Gases Section 3 Review Answers Budget-Friendly Options
6. Navigating Gases Section 3 Review Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Gases Section 3 Review Answers Compatibility with Devices
 - Gases Section 3 Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Gases Section 3 Review Answers
 - Highlighting and Note-Taking Gases Section 3 Review Answers
 - Interactive Elements Gases Section 3 Review Answers
8. Staying Engaged with Gases Section 3 Review Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gases Section 3 Review Answers
9. Balancing eBooks and Physical Books Gases Section 3 Review Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gases Section 3 Review Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gases Section 3 Review Answers
- Setting Reading Goals Gases Section 3 Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gases Section 3 Review Answers
- Fact-Checking eBook Content of Gases Section 3 Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

-
- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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