



# Covalent Compounds Unit Study Guide

Name \_\_\_\_\_

Teacher \_\_\_\_\_

Date: \_\_\_\_\_

Period \_\_\_\_\_

## Reference:

Chemistry in the Community, pp 181-183,  
Notes  
Naming covalent compounds worksheet  
Viscosity Lab  
Exploring Properties of Ionic and Covalent compounds

## Vocabulary

Understand and apply the definitions of the following terms:

Covalent

Valence electrons

e dot diagrams (Lewis dot diagrams)

Polar

Nonpolar

Intermolecular forces

Octet rule

Polymers

Viscosity

Boiling points

## Targeted Skills

- Diagram a covalent molecule using a e dot diagrams (Lewis dot diagrams).
- Know the following properties:
  - > Ionic compounds = the transfer of electron(s), i.e. "gain or lose"  
Metal cations & nonmetal anion (or polyatomic ion)  
Net charge = zero (criss cross charges)  $X^{+n}Y^{-m} = X_nY_m$   
Dissociate (dissolve) in water
  - > Molecular compounds = sharing electrons by forming a covalent bond  
2 or more nonmetals  
Prefixes indicate # of each atom  
Polar molecules have a  $\delta^+$  & a  $\delta^-$   
Polar molecules dissolve in water  
Nonpolar molecules do not have oppositely charged ends, therefore do not dissolve in water.
  - > Octet rule = atoms share, gain, or lose e- to obtain full outer energy level (usually 8 e-)
- Name covalent compounds using prefixes: (mono, di, tri, tetra, penta, hexa, hepta, octa, nona, deca).
- Compare/contrast the arrangement of atoms in molecules to the arrangement of atoms in ionic compounds.
- Describe the influence of intermolecular forces on physical and chemical properties of covalent compounds. Example: Boiling points, viscosity, and density.
- Be able to describe properties (boiling points, melting points, viscosity, solubility, conductivity) of polymers, ionic crystals, metallic substances, atoms and molecules.

## Make sure you remember from past tests?

Atomic structure  
Significant figures  
Accuracy & precision  
Mixtures  
Nomenclature

Solubility  
Safety  
Elements  
Ionic compound formula writing

# Holt Chemistry Covalent Compounds Study Guide

**Guillaume Favre**



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## **Table of Contents Holt Chemistry Covalent Compounds Study Guide**

1. Understanding the eBook Holt Chemistry Covalent Compounds Study Guide
  - The Rise of Digital Reading Holt Chemistry Covalent Compounds Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Holt Chemistry Covalent Compounds Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Holt Chemistry Covalent Compounds Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Chemistry Covalent Compounds Study Guide
  - Personalized Recommendations
  - Holt Chemistry Covalent Compounds Study Guide User Reviews and Ratings

- Holt Chemistry Covalent Compounds Study Guide and Bestseller Lists
- 5. Accessing Holt Chemistry Covalent Compounds Study Guide Free and Paid eBooks
  - Holt Chemistry Covalent Compounds Study Guide Public Domain eBooks
  - Holt Chemistry Covalent Compounds Study Guide eBook Subscription Services
  - Holt Chemistry Covalent Compounds Study Guide Budget-Friendly Options
- 6. Navigating Holt Chemistry Covalent Compounds Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Holt Chemistry Covalent Compounds Study Guide Compatibility with Devices
  - Holt Chemistry Covalent Compounds Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Holt Chemistry Covalent Compounds Study Guide
  - Highlighting and Note-Taking Holt Chemistry Covalent Compounds Study Guide
  - Interactive Elements Holt Chemistry Covalent Compounds Study Guide
- 8. Staying Engaged with Holt Chemistry Covalent Compounds Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Holt Chemistry Covalent Compounds Study Guide
- 9. Balancing eBooks and Physical Books Holt Chemistry Covalent Compounds Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Holt Chemistry Covalent Compounds Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Holt Chemistry Covalent Compounds Study Guide
  - Setting Reading Goals Holt Chemistry Covalent Compounds Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Holt Chemistry Covalent Compounds Study Guide
  - Fact-Checking eBook Content of Holt Chemistry Covalent Compounds Study Guide
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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