

Chem 104A Study Guide

1. A gas is initially at 200 kPa pressure and 27°C. If the temperature is increased by 10 degrees Celsius, what will the new pressure be if volume remains constant?
2. The volume of a gas is 1.1 L and at 400 kPa. Assuming ideal behavior, what is the volume at the temperature of the gas?
3. A mixture that contained equal moles of O_2 , N_2 , CO_2 , and CO (all at 20°C) was put into a graduated glass container containing 2.0 g of O_2 at 20°C and 1.0 atm pressure.

 - a. How many moles of gas are present and is it above or below the temperature of 20°C? What will be the new pressure of the gas if it expands the volume of the gas to 2.0 L and the pressure is 1.0 atm?
 - b. How many moles of gas are present and is it above or below the temperature of 20°C? What will be the new pressure of the gas if it expands the volume of the gas to 2.0 L and the pressure is 1.0 atm?

4. How many moles of gas are present and is it above or below the temperature of 20°C? What will be the new pressure of the gas if it expands the volume of the gas to 2.0 L and the pressure is 1.0 atm?

Gas Law Study Guide

James E. Brady, John R. Holum



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