

ICE Practice Problems

#1 – Relatively easy, no ICE table required because eq'm concentrations are given

For the reaction $\text{CH}_4(\text{g}) + \text{H}_2\text{O}(\text{g}) \leftrightarrow \text{CO}(\text{g}) + 3\text{H}_2(\text{g})$ @ 1500°C an equilibrium mixture of these gases was found to have the following concentrations $[\text{CO}] = 0.300\text{M}$, $[\text{H}_2] = 0.800\text{M}$ and $[\text{CH}_4] = 0.400\text{M}$. K_c @ $1500^\circ\text{C} = 5.67$. Determine the equilibrium concentration of H_2O in this mixture.

#2 – Requires an ICE table because you do not know the equilibrium concentrations- no product is yet formed

For the reaction $\text{CO}(\text{g}) + \text{H}_2\text{O}(\text{g}) \leftrightarrow \text{CO}_2(\text{g}) + \text{H}_2(\text{g})$ calculate the equilibrium concentrations of all species if 1.000 mol of each reactant is mixed in a 1.000L flask. $K_c = 5.10$ at the temperature of this reaction.

		$\text{CO}(\text{g})$	+	$\text{H}_2\text{O}(\text{g})$	\leftrightarrow	$\text{CO}_2(\text{g})$	+	$\text{H}_2(\text{g})$
[Initial]	I	1.000		1.000		0		0
[Change in]	C							
[Equilibrium]	E							

#3 – Requires an ICE table because you do not know the equilibrium concentrations. The initial concentrations must be calculated- no product is yet formed

For the reaction $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \leftrightarrow 2\text{HF}(\text{g})$ calculate the equilibrium concentrations of all species if 3.000 mol of each reactant was added 1.500L flask. K_c at the temperature of the reaction is 1.15×10^2 .

		$\text{H}_2(\text{g})$	+	$\text{F}_2(\text{g})$	\leftrightarrow	$2\text{HF}(\text{g})$
[Initial]	I	2.00		2.00		0
[Change in]	C					
[Equilibrium]	E					

#4 – Requires an ICE table because you do not know the equilibrium concentrations. Initial concentrations of reactants are given.

0.200mol of H_2 , 0.200mol of I_2 , and 0.200mol of HI were placed in a 1.00 L flask and allowed to come to equilibrium. The K_c value of the reaction at this temperature is 49.5. Determine the equilibrium concentrations of all species.

		$\text{H}_2(\text{g})$	+	$\text{I}_2(\text{g})$	\leftrightarrow	$2\text{HI}(\text{g})$
[Initial]	I					
[Change in]	C					
[Equilibrium]	E					

#5 – Requires an ICE table because you do not know the equilibrium concentrations. Initial concentrations of reactants must be calculated and no product is yet formed

For the reaction $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \rightarrow 2\text{HF}(\text{g})$ calculate the equilibrium concentrations of each species if 3.000 mol of H_2 and 6.000mol of F_2 are mixed in a 3.000L flask. K_c at this temperature is 1.15×10^2 .

		$\text{H}_2(\text{g})$	+	$\text{F}_2(\text{g})$	\leftrightarrow	$2\text{HF}(\text{g})$
[Initial]	I					
[Change in]	C					
[Equilibrium]	E					

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