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ENGINEERING ECONOMY

Sixteenth Edition

Engineering Economy Canadian Edition Solution Manual

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lithium li has three electrons two fill the 1 s orbital and the third is placed in the 2 s orbital giving an electron configuration of $1s^2 2s^1$

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are he 2 helium $1s^2$ ne 10 neon $1s^2 2s^2 2p^6$

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