

BUILD Vocabulary

ionic bond Chemical bond formed when one or more electrons are transferred from one atom to another.

ion atom that has a positive or negative charge.

covalent bond type of bond between atoms in which the electrons are shared.

molecule smallest unit of most compounds that displays all the properties of that compound.

van der Waals forces weak attraction that develops between oppositely charged regions of nearby molecules.

Preface The preface can be read on your own “together.” A compound is made up of elements joined together, and a covalent bond connects atoms that share electrons together. **How are covalent bonds different from ionic bonds?**

Chemical Bonds

KEY QUESTION What are the main types of chemical bonds?

The atoms in compounds are held together by chemical bonds. Bonds are formed by the electrons in the outer shell, which are called valence electrons. The main types of chemical bonds are ionic bonds and covalent bonds.

Ionic Bonds **ionic bonds** are formed when one or more electrons are transferred from one atom to another. This transfer causes the atom losing the electrons to become positively charged. The atom gaining the electrons becomes negatively charged. The positively and negatively charged atoms become **ions**. The oppositely charged ions are attracted to each other and form an ionic bond.

Covalent Bonds Sometimes atoms share electrons rather than transferring them. These shared electrons move around the nuclei of both atoms. The bond formed by sharing electrons is called a **covalent bond**. Atoms can share between two and six electrons. The structure that results is known as a **molecule**. A molecule is the smallest unit of most compounds.

Weak Interactions The strongest chemical bonds are ionic or covalent. Weak interactions between atoms are also important. Within living organisms, many molecules interact only briefly. **Van der Waals forces** are one type of weak interaction between molecules. They occur when molecules are very close together.

Hydrogen bonds are another type of weak interaction. They occur between a hydrogen atom of one molecule and an oxygen or nitrogen atom of a neighboring molecule.

Visualizing with Covalent Bonds

1. What kind of chemical bond holds the water molecule together? _____
2. What particles are shared between the atoms in the molecule? _____
3. What is the other main type of chemical bond? _____
4. What is the difference between covalent and ionic bonds? _____



Water molecule (H₂O)

READING LOOK Cause and Effect As you read the part of the lesson under “The Water Molecule,” use the table below to list the causes and effects of the properties of water. Many of the causes may be the same, but each characteristically will have a different effect. Part of the table is completed for you.

Characteristics of Water	Cause	Effect
Hydrogen Bonding		The slightly negative end of a water molecule is attracted to the slightly positive end of another water molecule.
Expands Upon Freezing		
Dissolves Many Substances	Water forms hydrogen bonds.	
Cohesion		
Adhesion		
Heat Capacity		

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Table of Contents Miller And Levine Chapter 2 Workbook

1. Understanding the eBook Miller And Levine Chapter 2 Workbook
 - The Rise of Digital Reading Miller And Levine Chapter 2 Workbook
 - Advantages of eBooks Over Traditional Books
2. Identifying Miller And Levine Chapter 2 Workbook
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Miller And Levine Chapter 2 Workbook
 - User-Friendly Interface
4. Exploring eBook Recommendations from Miller And Levine Chapter 2 Workbook

- Personalized Recommendations
 - Miller And Levine Chapter 2 Workbook User Reviews and Ratings
 - Miller And Levine Chapter 2 Workbook and Bestseller Lists
5. Accessing Miller And Levine Chapter 2 Workbook Free and Paid eBooks
 - Miller And Levine Chapter 2 Workbook Public Domain eBooks
 - Miller And Levine Chapter 2 Workbook eBook Subscription Services
 - Miller And Levine Chapter 2 Workbook Budget-Friendly Options
 6. Navigating Miller And Levine Chapter 2 Workbook eBook Formats
 - ePub, PDF, MOBI, and More
 - Miller And Levine Chapter 2 Workbook Compatibility with Devices
 - Miller And Levine Chapter 2 Workbook Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Miller And Levine Chapter 2 Workbook
 - Highlighting and Note-Taking Miller And Levine Chapter 2 Workbook
 - Interactive Elements Miller And Levine Chapter 2 Workbook
 8. Staying Engaged with Miller And Levine Chapter 2 Workbook
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Miller And Levine Chapter 2 Workbook
 9. Balancing eBooks and Physical Books Miller And Levine Chapter 2 Workbook
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Miller And Levine Chapter 2 Workbook
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Miller And Levine Chapter 2 Workbook
 - Setting Reading Goals Miller And Levine Chapter 2 Workbook
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Miller And Levine Chapter 2 Workbook

- Fact-Checking eBook Content of Miller And Levine Chapter 2 Workbook
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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