

Worksheet # 1 & Gram Formula Mass Answers

- | | |
|--|------------------|
| 1. 4.5×10^{23} mcls H_2O | 1. 158.04 g/mol |
| 2. 3×10^{24} atoms Fe | 2. 74.55 g/mol |
| 3. 5.73×10^{-4} mol Cs | 3. 142.05 g/mol |
| 4. 1.1×10^5 mol Na | 4. 164.10 g/mol |
| 5. 2.1×10^{23} mcls $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ | 5. 342.17 g/mol |
| 6. 9.55×10^{-2} mol Cu | 6. 149.12 g/mol |
| 7. 1.204×10^{24} atoms Fr | 7. 249.70 g/mol |
| 8. 2.000 mol He | 8. 262.87 g/mol |
| 9. 5.6×10^2 mol S | 9. 219.52 g/mol |
| | 10. 458.16 g/mol |

Gram Formula Mass Answers

Joseph A. Mascetta



Gram Formula Mass Answers:

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section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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Modern Chemistry, 1990

Chemistry John McMurry, 1998

English Mechanic and Mirror of Science and Art, 1918

Introduction to Chemical Principles Howard Stephen Stoker, 1990 Introduction to Chemical Principles is a text for students who have had little to no previous instruction in chemistry or who had such instruction long enough ago that a thorough review is needed

preface

Preparatory Chemistry Howard Stephen Stoker, 1990

A Standardized Test in the Mechanics of Physics Harriet Alma Ranney, 1928

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