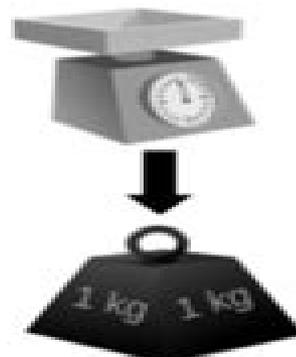


Intro to Stoichiometry – Grams to Grams NOTES

Step-by-Step

If you have grams of a chemical and want grams of a different one

- 1) Convert from grams to moles.
- 2) Convert moles of one chemical into moles of another chemical.
- 3) Convert moles of your NEW chemical into grams of that chemical.



Ex. 1: Given 68 grams of H_2 , how many grams of O_2 are produced?

$$\frac{68 \text{ grams H}_2}{1} \times \frac{1 \text{ mole H}_2}{2.02 \text{ grams H}_2} \times \frac{1 \text{ mole O}_2}{2 \text{ moles H}_2} \times \frac{32.00 \text{ grams O}_2}{1 \text{ mole O}_2}$$

Answer: 538.61 grams of O_2

Ex. 2: Given 2 grams of O_2 , how many grams of H_2 are produced?

$$\frac{2 \text{ grams O}_2}{1} \times \frac{1 \text{ mole O}_2}{32.00 \text{ grams O}_2} \times \frac{2 \text{ moles H}_2}{1 \text{ mole O}_2} \times \frac{2.02 \text{ grams H}_2}{1 \text{ mole H}_2}$$

Answer: 0.25 grams of H_2

Ex. 3: Given 8.35 grams of H_2 , how many grams of H_2O are produced?

$$\frac{8.35 \text{ grams H}_2}{1} \times \frac{1 \text{ mole H}_2}{2.02 \text{ grams H}_2} \times \frac{2 \text{ moles H}_2\text{O}}{2 \text{ moles H}_2} \times \frac{18.02 \text{ grams H}_2\text{O}}{1 \text{ mole H}_2\text{O}}$$

Answer: 74.49 grams of H_2O

Ex. 4: Given 0.94 grams of H_2 , how many gram of O_2 are produced?

$$\frac{0.94 \text{ grams H}_2}{1} \times \frac{1 \text{ mole H}_2}{2.02 \text{ grams H}_2} \times \frac{1 \text{ mole O}_2}{2 \text{ moles H}_2} \times \frac{32.00 \text{ grams O}_2}{1 \text{ mole O}_2}$$

Answer: 7.45 grams of O_2

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