

Holt Modern Chemistry Review

CHAPTER 6: CHEMICAL BONDING

The following pages contain the bulk (but not all) of the information for the chapter 6 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and NOTHING else, you should at least be able to PASS the test. ***** Test items will be recall, examples, and/or application of this content. *****

OUTCOMES

- Collaborate with peer(s) to understand chemistry content (C C)
- Communicate chemistry content to teacher and peer(s) (E C)
- 6.2: Use the Lewis dot notation, illustrate the covalent bond as the sharing of electrons (F & PK)
- 6.3: List important physical properties of ionic compounds; sum of oxidation numbers = zero (F & PK)

6.1: INTRODUCTION TO CHEMICAL BONDING

- **Vocabulary**
 - **chemical bond** -- the attractive force that holds atoms or ions together
 - **ionic bonding** -- a force that attracts electrons from one atom to another, which transforms a neutral atom into an ion
 - **covalent bonding** -- a bond formed when atoms share one or more pairs of electrons
 - **nonpolar-covalent bond** -- a covalent bond in which the bonding electrons are equally attracted to both bonded atoms
 - **polar** -- describes a molecule in which the positive and negative charges are separated
 - **polar-covalent bond** -- a covalent bond in which a pair of electrons shared by two atoms is held more closely by one atom
- **Chapter Highlights**
 - Most atoms are chemically bonded to other atoms. The three major types of chemical bonding are ionic, covalent, and metallic.
 - In general, atoms of metals bond ionically with atoms of nonmetals, atoms of metals bond metallically with each other, and atoms of nonmetals bond covalently with each other.

6.2: COVALENT BONDING AND MOLECULAR COMPOUNDS

- **Vocabulary**
 - **molecule** -- a group of atoms that are held together by chemical forces; a molecule is the smallest unit of matter that can exist by itself and retain all of a substance's chemical properties
 - **molecular compound** -- a chemical compound whose simplest units are molecules
 - **chemical formula** -- a combination of chemical symbols and numbers to represent a substance
 - **molecular formula** -- a chemical formula that shows the number and kinds of atoms in a molecule, but not the arrangement of the atoms
 - **electron-dot notation** -- an electron configuration notation in which only the valence electrons of an atom of a particular element are shown, indicated by dots placed around the element's symbol
 - **Lewis structure** -- a structural formula in which electrons are represented by dots; dot pairs or dashes between two atomic symbols represent pairs in covalent bonds.
See examples at right for LEWIS STRUCTURES of NF₃ and OCl₂
 - **structural formula** -- a formula that indicates the location of the atoms, groups, or ions relative to one another in a molecule and that indicates the number and location of chemical bonds
 - **single bond** -- a covalent bond in which two atoms share one pair of electrons
 - **multiple bond** -- a bond in which the atoms share more than one pair of electrons, such as a double bond or a triple bond
 - **resonance** -- the bonding in molecules or ions that cannot be correctly represented by a single Lewis structure
- **Chapter Highlights**
 - Atoms in molecules are joined by covalent bonds. In a covalent bond, two atoms share one or more pairs of electrons.



Holt Chemistry Chapter 6

David Kirk



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