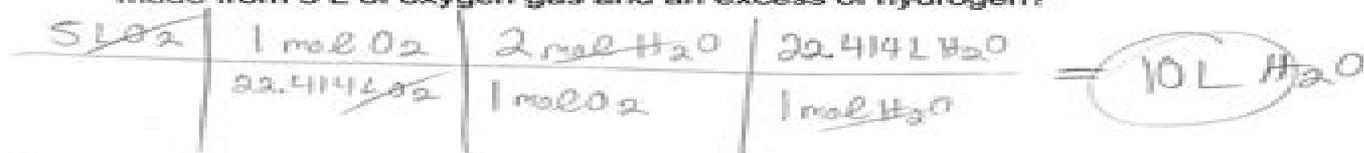


$$PV = nRT$$

0.0°C 1 atm
22.414 L

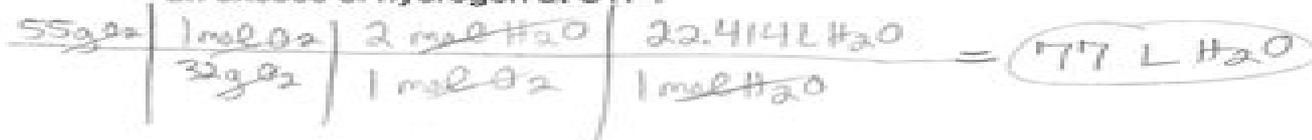
Gas Stoichiometry Practice Sheet

- 1) For the reaction $2 \text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2 \text{H}_2\text{O}_{(l)}$, how many liters of water can be made from 5 L of oxygen gas and an excess of hydrogen?

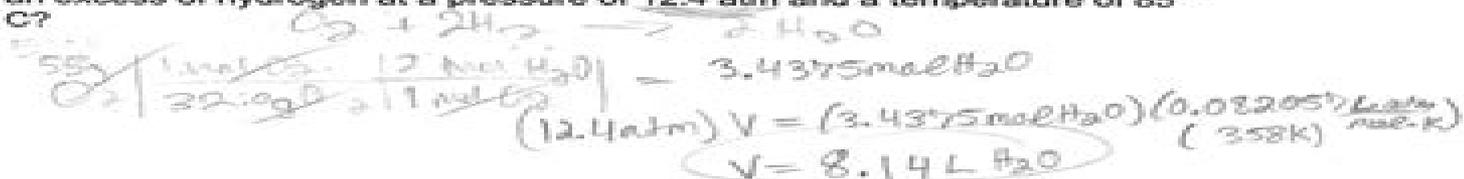


0.0°C
1 atm

- 2) How many liters of water can be made from 55 grams of oxygen gas and an excess of hydrogen at STP?



- 3) How many liters of water can be made from 55 grams of oxygen gas and an excess of hydrogen at a pressure of 12.4 atm and a temperature of 85°C?



- 4) How many liters of water can be made from 34 grams of oxygen gas and

4.5 L H₂O

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