

Final study guide question and answers

What kind of equilibrium exists and under what conditions? **Chemical equilibrium** is a dynamic equilibrium between two or more states of matter, where the concentrations of the reactants and products are constant over time.

What happens if you change the conditions? **Le Chatelier's principle** states that if a system at equilibrium is subjected to a change in concentration, temperature, or pressure, the system will adjust itself to counteract the change and restore a new equilibrium.

What happens if you change the volume of the container? **Volume change** affects the equilibrium position of a reaction involving gases. Increasing the volume shifts the equilibrium towards the side with more moles of gas, while decreasing the volume shifts it towards the side with fewer moles of gas.

What happens if you change the temperature? **Temperature change** affects the equilibrium position of a reaction. For an exothermic reaction, increasing the temperature shifts the equilibrium towards the reactants, while for an endothermic reaction, it shifts towards the products.

What happens if you change the concentration of the reactants or products? **Concentration change** affects the equilibrium position of a reaction. Increasing the concentration of a reactant shifts the equilibrium towards the products, while increasing the concentration of a product shifts it towards the reactants.

What happens if you change the pressure of the system? **Pressure change** affects the equilibrium position of a reaction involving gases. Increasing the pressure shifts the equilibrium towards the side with fewer moles of gas, while decreasing the pressure shifts it towards the side with more moles of gas.

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Mark Steyvers



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