

CONFORMS TO THE 2011 NEC®

McGRAW-HILL'S

National Electrical Code® **2011** Handbook

TWENTY-SEVENTH
EDITION

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of hcl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0 100 m the ph of the solution is then $\text{ph} = -\log 0.100 = 1.000$ ph log 0 100 1 000 ph log 0 100 1 000 b titrant volume 12 50 ml

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