

1. Electrochemistry Parts 1&2 - Alec, Daniela, Ian, Varena - 03/22/2023 - CHM2046L.019 - 1

2. Aim:

The aim of this experiment is to learn how to set up a galvanic cell and utilize it to test the electrode potentials of various metals in varying conditions. This leads to us constructing a galvanic cell that will help us measure the redox potential.

3. Materials:

- Gloves
- Goggles
- Lab pants (scrubs)
- Lab coat
- Closed-toed shoes
- 24- well plate
- Multimeter
- 5 ml Pipette
- 10 mL Graduated cylinder
- Sandpaper
- Electrodes: Cu, Zn, Pb Fe, C
- KCL, $\text{Cu}(\text{NO}_3)_2$, $\text{Pb}(\text{NO}_3)_2$, FeSO_4 or $\text{Fe}(\text{NO}_3)_3$

-Instruction to set up a voltaic cell

1. The cell basically consists of an anode(being oxidized), a cathode(being reduced), and a salt bridge.
2. Put the cathode and anode metal strips into the solutions that are going to be used during the experiment, and after that connected the solutions using the salt bridge.
3. Go to the multimeter, plug in the leads, and turn it on.
4. Choose the 20 dc volts setting.
5. Now, it is time to measure the voltage of the cell. Connect the red lead to the cathode and the black lead to the anode. The voltage measure will be displayed.

- Instructions to use a multimeter

1. Plug the black probe into the **COM** port on your multimeter
2. Plug the red probe into the **VΩmA** port.
3. Set the dial to continuity mode, which looks like a triangle with a line on the right side.
4. Connect the red lead, which is positive to the cathode, and the black lead to the anode terminal. Make sure you observe a positive voltage.

4. Chemicals/Hazards:

Electrochemical Lab Pre Lab Answers

Sabine Zange



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