

## EXERCISE # (O-I)

## Electrode Potential, Galvanic Cell

- In a galvanic cell
  - chemical reaction produces electrical energy
  - electrical energy produces chemical reaction
  - reduction occurs at anode
  - oxidation occurs at cathode
- Which of the following statement is wrong about galvanic cell ?
  - Cathode is positive charged
  - Anode is negatively charged
  - Reduction takes place at the anode
  - Electrons move from anode to cathode
- Which of the following has been universally accepted as a reference electrode at all temperatures and has been assigned a value of zero volt?
  - Standard platinum electrode
  - Standard copper electrode
  - Standard graphite electrode
  - Standard hydrogen electrode
- A standard hydrogen electrode has zero electrode potential because
  - hydrogen is easier to oxidise
  - electrode potential is assumed to be zero
  - hydrogen atom has only one electron
  - hydrogen is the lightest element.
- Which is not true for a standard hydrogen electrode ?
  - The hydrogen ion concentration must be 1 M
  - Temperature must be 25°C
  - Pressure of hydrogen must be 1 bar
  - Temperature can vary
- The equation representing the process by which standard reduction potential of zinc can be defined is :
  - $\text{Zn}^{2+}(\text{s}) + 2\text{e}^- \longrightarrow \text{Zn}(\text{s})$
  - $\text{Zn}(\text{g}) \longrightarrow \text{Zn}^{2+}(\text{g}) + 2\text{e}^-$
  - $\text{Zn}^{2+}(\text{g}) + 2\text{e}^- \longrightarrow \text{Zn}(\text{s})$
  - $\text{Zn}^{2+}(\text{aq.}) + 2\text{e}^- \longrightarrow \text{Zn}(\text{s})$

$\Delta G = -nFE$  theoretical
- $E^\circ$  for  $\text{F}_2 + 2\text{e}^- = 2\text{F}^-$  is 2.8 V,  $E^\circ$  for  $\frac{1}{2}\text{F}_2 + \text{e}^- = \text{F}^-$  is ?
  - 2.8 V
  - 1.4 V
  - 2.8 V
  - 1.4 V
- Doubling all the coefficients in the equation for the cell reaction -
  - Doubles both  $E^\circ$  and  $\Delta G^\circ$
  - Doubles  $E^\circ$ , but does not change  $\Delta G^\circ$
  - Doubles  $\Delta G^\circ$ , but does not change  $E^\circ$
  - Does not change  $E^\circ$  or  $\Delta G^\circ$

# Electrochemistry Exercise Answers

**Karin Nielsen-Saines**



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