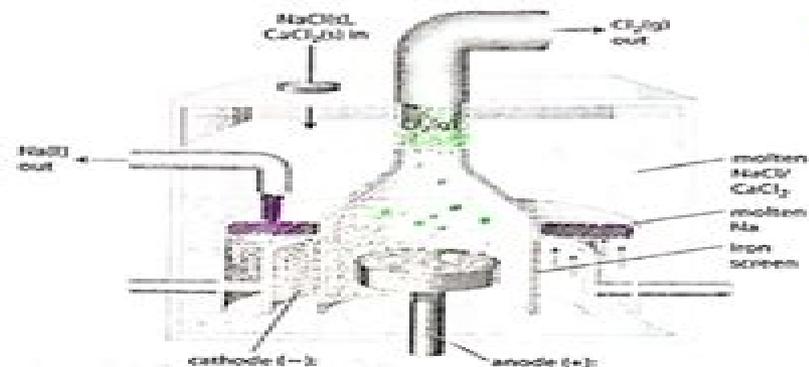


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Electrolysis of sodium chloride (in the industry)

Industrial electrolysis of molten Sodium Chloride requires extreme temperatures.

- Calcium Chloride is added to the mixture to lower the melting point of Sodium Chloride.
- Molten sodium metal floats in the mixture because of its low density.
- The anode is carbon
- The cathode is iron
- Chlorine gas is collected separately



3.

4. electrolysis of Aluminum oxide (molten)

Electrodes: graphite (carbon)

Electrolyte: molten Alumina (Al_2O_3)

Alumina = Purified Bauxite

Which substance is added to the molten alumina? cryolite

Ions Present: Al^{3+} , O^{2-}

Half equations:

In the Cathode $(Al^{3+} + 3e^- \rightarrow Al) \times 4$

In the Anode $(2O^{2-} \rightarrow O_2 + 4e^-) \times 3$

Overall Equation $4 Al^{3+} + 6 O^{2-} \rightarrow 4 Al + 3 O_2$

Observations:

Cathode: The Molten Aluminum can be observed

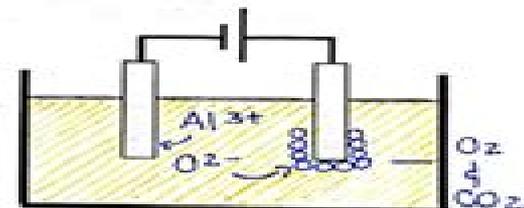
Anode: Bubbles of Oxygen gas will appear

Why is it necessary to change the electrodes periodically?

Because the anodes will burn out. Anode reacts with O_2 to form CO_2

Which is the purpose of the cryolite?

To lower the melting point of the Aluminum Oxide



Electrochemistry Answers Study Guide

Wolfgang Guggemos



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Basic Concepts in Electrochemistry Polagony Venkateshwarlu, 2024-09-30 As you know that electrochemistry is a fascinating subject that plays an important role in our everyday life activities it is the study of chemical processes that cause electrons to move This movement of electrons what we commonly known as electricity Powers many of the devices and systems that are integral to modern living This brilliant book provides a reference for students especially grade X and XII who want to learn basic principles in electrochemistry It is designed to provide fundamental principles of electrochemistry in simple manner which can be suitable to the needs of students and their better understanding of theoretical and practical applications in electrochemistry An attempt has been made to discuss about types of electrochemical cells on evolutionary basis from voltaic cell to sodium ion battery Simple outline has been covered based on structure components and functioning At the last a great attempt is also made to provide solution for NCERT exercise and some worksheet provided for practice of concepts

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