

Electrolysis Test Study Guide

What are the 8 uses of electrolysis? - Answer- Purification of metals

Electroplating

Chlorine production

Extraction of Hydrogen from water

Extraction of oxygen from water

Extraction of Reactive metals

Hair removal

Production of sodium hydroxide

What is Electrolysis? - Answer- The passing of an electric current through an molten ionic compound causing a chemical change resulting in a change in the elements that make up the compound.

What is an Electrolyte? - Answer- A molten ionic compound that undergoes electrolysis.

What is an Anode? - Answer- A positively charge electrode.

What is a Cathode? - Answer- a negatively charged electrode.

What is an electrode? - Answer- an electrical conductor.

(Chem web def: Terminals/sticks that allow electricity to go through them.)

What are the Diatomic elements? - Answer- H, N, O, F, Cl, Br, I

Or H-NOF-I

What Does being a Diatomic element mean? - Answer- This means that a when a comped is split up, and one of the products of the reaction is a diatomic element then that diatomic element has X_2

Tip: When a compound is split up, if it becomes a diatomic element and a metal, the diatomic element will be a gas (g), while the metal will be a solid (s).

What are the charges of the elements On the periodic table in Groups 1, 2, 3, 5, 6, 7 - Answer- Groups: 1 = +1, 2 = +2, 3 = +3, 5 = -3, 6 = -2, 7 = -1, (group 4 is neutral)

What are the steps for electrolysis? - Answer- 1. Draw the two electrodes and display their two charges.

a. Electrode is an electric conductor.

b. Their must be a positive (anode)and negative conductor (cathode)

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