

B. Chlorine

D. Calcium

F. Antimony

2. Draw Lewis Structures for each of the following elements. How many electrons are in their outer shells?

A. Silicon

C. Sulfur

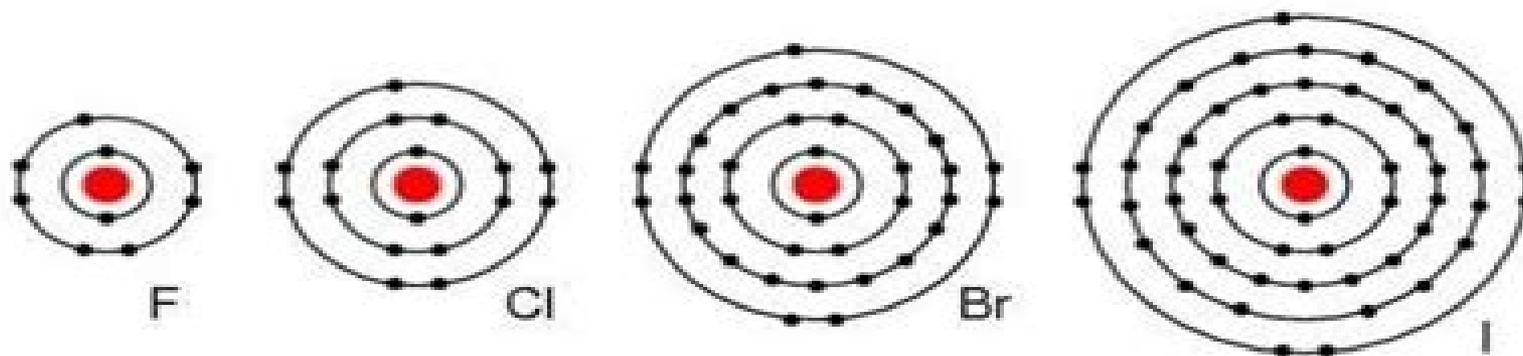
E. Beryllium

B. Potassium

D. Boron

F. Argon

3. Circle the valence electrons only. How many valence electrons does each atom have?



4. Finish the sentence: Elements in the same group have the same number of \_\_\_\_\_ electrons.

5. Fill in the electrons for the following elements. Use your work to answer how many valence electrons each element has.

Example: C (carbon): 1S  $\uparrow\downarrow$  2S  $\uparrow\downarrow$  2P  $\uparrow \uparrow$  \_\_\_\_\_ 3S \_\_\_\_\_ 3P \_\_\_\_\_ #valence electrons 4

He: 1S \_\_\_\_\_ 2S \_\_\_\_\_ 2P \_\_\_\_\_ 3S \_\_\_\_\_ 3P \_\_\_\_\_ Number of valence electrons in helium \_\_\_\_\_

Cl: 1S \_\_\_\_\_ 2S \_\_\_\_\_ 2P \_\_\_\_\_ 3S \_\_\_\_\_ 3P \_\_\_\_\_ Number of valence electrons in chlorine \_\_\_\_\_

Li: 1S \_\_\_\_\_ 2S \_\_\_\_\_ 2P \_\_\_\_\_ 3S \_\_\_\_\_ 3P \_\_\_\_\_ Number of valence electrons in lithium \_\_\_\_\_

B: 1S \_\_\_\_\_ 2S \_\_\_\_\_ 2P \_\_\_\_\_ 3S \_\_\_\_\_ 3P \_\_\_\_\_ Number of valence electrons in boron \_\_\_\_\_

F: 1S \_\_\_\_\_ 2S \_\_\_\_\_ 2P \_\_\_\_\_ 3S \_\_\_\_\_ 3P \_\_\_\_\_ Number of valence electrons in fluorine \_\_\_\_\_

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