

## Electron Configurations - Solutions

*Note: The electron configurations in this worksheet assume that lanthanum (La) is the first element in the 4f block and that actinium (Ac) is the first element in the 5f block. If your periodic table doesn't agree with this, your answers for elements near the f-orbitals may be slightly different.*

1)	sodium	$1s^2 2s^2 2p^6 3s^1$
2)	iron	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
3)	bromine	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
4)	barium	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
5)	thallium	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^1$
6)	cobalt	$[\text{Ar}] 4s^2 3d^7$
7)	silver	$[\text{Kr}] 5s^2 4d^9$
8)	tellurium	$[\text{Kr}] 5s^2 4d^{10} 5p^4$
9)	radium	$[\text{Rn}] 7s^2$
10)	astatine	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^5$
11)	lithium	$1s^2 2s^1$
12)	oxygen	$1s^2 2s^2 2p^4$
13)	calcium	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$
14)	titanium	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^2$
15)	rubidium	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$
16)	lead	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^8$
17)	bismuth	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^3$
18)	chlorine	$1s^2 2s^2 2p^6 3s^2 3p^5$
19)	iron	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
20)	zinc	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10}$
21)	polonium	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^4$
22)	tellurium	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^4$
23)	tungsten	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^4$
24)	molybdenum	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^4$

# Electron Configuration Practice Packet

**Kathy A. Zahler**



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