

8 • Electron Configurations & Periodicity

WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
Element: Ar # of e ⁻ 's: <u>18</u>	Element: Mg # of e ⁻ 's: <u>12</u>	Element: N # of e ⁻ 's: <u>7</u>	Element: Li # of e ⁻ 's: <u>3</u>	Element: P # of e ⁻ 's: <u>15</u>	Element: Cl # of e ⁻ 's: <u>17</u>

Write the electron configurations of each of these in **long form** and **short form**:

- Ar 1s² 2s² 2p⁶ 3s² 3p⁶
 Ar [Ne] 3s² 3p⁶
- Mg 1s² 2s² 2p⁶ 3s²
 Mg [Ne] 3s²
- N 1s² 2s² 2p³
 N [He] 2s² 2p³
- Li 1s² 2s¹
 Li [He] 2s¹
- P 1s² 2s² 2p⁶ 3s² 3p³
 P [Ne] 3s² 3p³
- Cl 1s² 2s² 2p⁶ 3s² 3p⁵
 Cl [Ne] 3s² 3p⁵



7. Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.



Electron Distributions Review Answer Key

Michael Brown



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