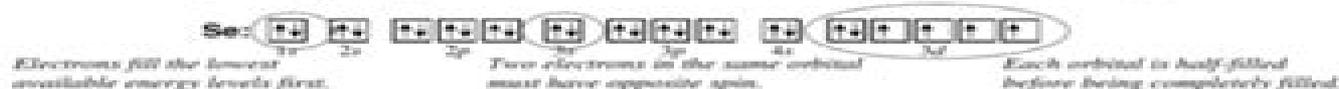


# Orbital Diagrams

## Chem Worksheet 5-5

Name \_\_\_\_\_

An orbital diagram uses boxes with arrows to represent the electrons in an atom. Each box in an orbital diagram represents an orbital. Orbitals have a capacity of two electrons. Arrows are drawn inside the boxes to represent electrons. Two electrons in the same orbital must have opposite spin so the arrows are drawn pointing in opposite directions. The following is an orbital diagram for selenium.

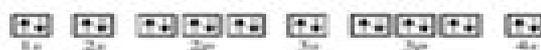
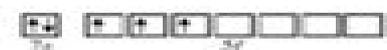


In writing an orbital diagram the first step is to determine the number of electrons. Normally this is the same as the number of protons, which is known as the atomic number. Next the boxes are drawn for the orbitals. Arrows are drawn in the boxes starting from the lowest energy sublevel and working up. This is known as the **Aufbau rule**. The **Pauli exclusion principle** requires that electrons in the same orbital have opposite spin. **Hund's rule** states that orbitals in a given sublevel are half-filled before they are completely filled.

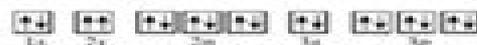
Boxes drawn for various sublevels		
<i>s</i> sublevel:	□	1 orbital
<i>p</i> sublevel:	□ □ □	3 orbitals
<i>d</i> sublevel:	□ □ □ □ □	5 orbitals
<i>f</i> sublevel:	□ □ □ □ □ □ □	7 orbitals



Write the name and symbol for the elements with the following orbital diagrams.

- |   |  |
|---|--|
| <p>1. </p> <p>2. </p> <p>3. </p> | <p>4. </p> <p>5. [Kr] </p> <p>6. [Rn] </p> |
|---|--|

There is an error with each of the following orbital diagrams. Explain the error.

- |   |  |
|---|--|
| <p>7. [Ar] </p> | <p>8. </p> |
|---|--|

Write orbital diagrams for the following. You may abbreviate using a noble gas.

- |  |   |
|--|---|
| <p>9. hydrogen</p> <p>10. boron</p> <p>11. sodium</p> <p>12. krypton</p> <p>13. chromium</p> <p>14. phosphorus</p> | <p>15. carbon</p> <p>16. cobalt</p> <p>17. platinum</p> <p>18. plutonium</p> <p>19. oxygen</p> <p>20. potassium</p> |
|--|---|

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**M Woodhall**



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