



# ELECTRON CONFIGURATIONS OF ATOMS AND IONS

## OBJECTIVES

- determine electron configuration of various elements/ions
- write electron configurations for various atoms, molecules, ions and ions
- determine quantum numbers for various elements as well as molecules, ions and ions

## THEORY

Each atom consists of protons and neutrons in nucleus and electrons orbiting a central nucleus. The electrons orbiting nucleus are arranged in shells or energy levels. The electrons in the outermost shell are called valence electrons. The electrons in the inner shells are called core electrons.

## EXERCISES

1. Write the electron configuration of the following elements and ions. Also write the quantum numbers of each electron.

Li (Z=3)	Be (Z=4)	B (Z=5)
C (Z=6)	N (Z=7)	O (Z=8)
F (Z=9)	Ne (Z=10)	Na (Z=11)
Mg (Z=12)	Al (Z=13)	Si (Z=14)
P (Z=15)	S (Z=16)	Cl (Z=17)
Ar (Z=18)	K (Z=19)	Ca (Z=20)
Sc (Z=21)	Ti (Z=22)	V (Z=23)
Cr (Z=24)	Mn (Z=25)	Fe (Z=26)
Co (Z=27)	Ni (Z=28)	Cu (Z=29)
Zn (Z=30)	Ga (Z=31)	Ge (Z=32)
As (Z=33)	Se (Z=34)	Br (Z=35)
Kr (Z=36)	Rb (Z=37)	Sr (Z=38)
Y (Z=39)	Zr (Z=40)	Nb (Z=41)
Mo (Z=42)	Tc (Z=43)	Ru (Z=44)
Rh (Z=45)	Pd (Z=46)	Ag (Z=47)
Au (Z=79)	Hg (Z=80)	Tl (Z=81)
Pb (Z=82)	Bi (Z=83)	Po (Z=84)
At (Z=85)	Rn (Z=86)	Fr (Z=87)
Ra (Z=88)	Ac (Z=89)	Th (Z=90)
Pa (Z=91)	U (Z=92)	Np (Z=93)
Pu (Z=94)	Am (Z=95)	Cm (Z=96)
Bk (Z=97)	Cf (Z=98)	Es (Z=99)
Fm (Z=100)	Md (Z=101)	No (Z=102)
Lr (Z=103)	Rf (Z=104)	Db (Z=105)
Sg (Z=106)	Bh (Z=107)	Hs (Z=108)
Mt (Z=109)	Ds (Z=110)	Rg (Z=111)

2. A particular element has the following quantum numbers:  $n=4, l=2, m=0, s=+\frac{1}{2}$ . Write the electron configuration of this element and name the element. Also write the quantum numbers of the other electrons in the atom.

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C (Z=6)	N (Z=7)	O (Z=8)
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# Electron Configurations Of Atoms And Ions Small Scale Experiment

**Ratko K. Janev**



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