

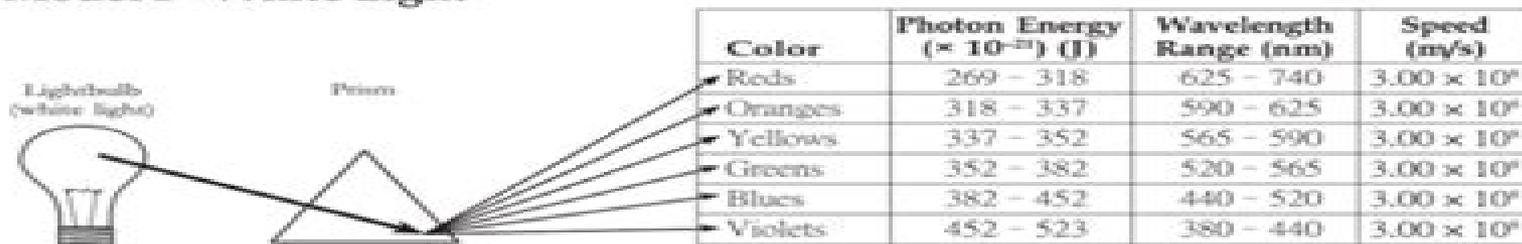
## Electron Energy and Light (Electron Excitation)

### Why?

How does light reveal the behavior of electrons in an atom?

From fireworks to stars, the color of light is useful in finding out what's in matter. The emission of light by hydrogen and other atoms has played a key role in understanding the electronic structure of atoms. Trace materials, such as evidence from a crime scene, lead in paint or mercury in drinking water, can be identified by heating or burning the materials and examining the color(s) of light given off in the form of bright-line spectra.

### Model 1 – White Light



1. Trace the arrows in Model 1 and shade in the table with colored pencils where appropriate.
2. What happens to white light when it passes through a prism?
  
3. Why are the color labels in the table in Model 1 plural (i.e., “Reds” rather than “Red”)?
  
4. Do all colors of light travel at the same speed?
  
5. Do all colors of light have the same energy? If no, which colors have the highest energy and the least energy, respectively?
  
6. Consider the light illustrated in Model 1.
  - a. Which color corresponds to the longest wavelengths?
  
  - b. Which color corresponds to the shortest wavelengths?
  
  - c. Write a sentence that describes the relationship between wavelength and energy of light.



# Electron Energy And Light Worsheet

**JE Gale**



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