

8 • Electron Configurations & Periodicity

ANSWERS TO STUDY QUESTIONS

- Write the electron configurations of the following elements using the shorthand notation for the noble gas cores.
 - phosphorus [Ne] $3s^2 3p^3$
 - nickel [Ar] $3d^8 4s^2$
 - osmium [Xe] $4f^{14} 5d^6 6s^2$
 - californium [Rn] $5f^{10} 7s^2$
 - titanium [Ar] $3d^2 4s^2$
- Which orbital is filled following these orbitals?
 - 3d is followed by 4p
 - 4s is followed by 3d
 - 5p is followed by 6s
 - 5f is followed by 6d
- How many electrons can be accommodated in
 - a d subshell 10; two in each of 5 orbitals
 - a set of f orbitals 14; two in each of 7 orbitals
 - the $n = 4$ shell 32; $2s^2$
 - the 7s orbital 2
 - a p_x orbital? 2
- What is wrong with the following ground state electron configurations?

a. $\begin{array}{c} 3d \\ \boxed{} \boxed{} \boxed{} \boxed{} \end{array}$ $\begin{array}{c} 4s \\ \boxed{1} \end{array}$ $\begin{array}{c} 4p \\ \boxed{} \boxed{} \end{array}$	a. OK, 4s is filled before 3d.
b. $\begin{array}{c} 3d \\ \boxed{1} \boxed{1} \boxed{1} \boxed{1} \boxed{1} \end{array}$ $\begin{array}{c} 4s \\ \boxed{1} \end{array}$ $\begin{array}{c} 4p \\ \boxed{1} \boxed{1} \boxed{1} \end{array}$	b. 4s should be filled before 4p subshell.
c. $\begin{array}{c} 3d \\ \boxed{1} \boxed{1} \boxed{1} \boxed{1} \boxed{1} \end{array}$ $\begin{array}{c} 4s \\ \boxed{1} \end{array}$ $\begin{array}{c} 4p \\ \boxed{1} \boxed{1} \boxed{1} \end{array}$	c. 4s cannot have 2 electrons with same spin.
d. $\begin{array}{c} 3d \\ \boxed{1} \boxed{1} \boxed{1} \boxed{1} \boxed{1} \end{array}$ $\begin{array}{c} 4s \\ \boxed{1} \end{array}$ $\begin{array}{c} 4p \\ \boxed{} \boxed{} \end{array}$	d. OK, stability of half-filled 3d subshell.
e. $\begin{array}{c} 3d \\ \boxed{1} \boxed{1} \boxed{1} \boxed{1} \boxed{1} \end{array}$ $\begin{array}{c} 4s \\ \boxed{1} \end{array}$ $\begin{array}{c} 4p \\ \boxed{1} \boxed{} \end{array}$	e. 4p should first fill one orbital at a time.

- How many unpaired electrons are there in
 - a nitrogen atom 3 unpaired electrons ($2s^2 2p^3$)
 - an iodine atom 1 unpaired electrons ($5s^2 5p^3$)
 - a nickel (II) cation 2 unpaired electrons ($3d^8$)
 - an oxide ion no unpaired electrons (closed shell [Ne])

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Structure, Bonding and the Periodic Law The Open University,Open University. Faculty of Science,Open University. Structure, Bonding and the Periodic Law Course Team,1973

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Table of Contents Electrons And Periodicity Packet Answers Unit 3

1. Understanding the eBook Electrons And Periodicity Packet Answers Unit 3
 - The Rise of Digital Reading Electrons And Periodicity Packet Answers Unit 3
 - Advantages of eBooks Over Traditional Books
2. Identifying Electrons And Periodicity Packet Answers Unit 3
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electrons And Periodicity Packet Answers Unit 3
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electrons And Periodicity Packet Answers Unit 3
 - Personalized Recommendations
 - Electrons And Periodicity Packet Answers Unit 3 User Reviews and Ratings

- Electrons And Periodicity Packet Answers Unit 3 and Bestseller Lists
- 5. Accessing Electrons And Periodicity Packet Answers Unit 3 Free and Paid eBooks
 - Electrons And Periodicity Packet Answers Unit 3 Public Domain eBooks
 - Electrons And Periodicity Packet Answers Unit 3 eBook Subscription Services
 - Electrons And Periodicity Packet Answers Unit 3 Budget-Friendly Options
- 6. Navigating Electrons And Periodicity Packet Answers Unit 3 eBook Formats
 - ePub, PDF, MOBI, and More
 - Electrons And Periodicity Packet Answers Unit 3 Compatibility with Devices
 - Electrons And Periodicity Packet Answers Unit 3 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electrons And Periodicity Packet Answers Unit 3
 - Highlighting and Note-Taking Electrons And Periodicity Packet Answers Unit 3
 - Interactive Elements Electrons And Periodicity Packet Answers Unit 3
- 8. Staying Engaged with Electrons And Periodicity Packet Answers Unit 3
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electrons And Periodicity Packet Answers Unit 3
- 9. Balancing eBooks and Physical Books Electrons And Periodicity Packet Answers Unit 3
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electrons And Periodicity Packet Answers Unit 3
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electrons And Periodicity Packet Answers Unit 3
 - Setting Reading Goals Electrons And Periodicity Packet Answers Unit 3
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electrons And Periodicity Packet Answers Unit 3
 - Fact-Checking eBook Content of Electrons And Periodicity Packet Answers Unit 3
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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