

## ANSWER KEY

# ATOMS AND ELECTRONS STUDY GUIDE

**Learning Goals:** (1) (10-1) P. 41 - Explain a new scientific theory of atomic-scale interactions (atomic theory) by describing changes in the atomic model over time, including those changes that are associated with experimental evidence. (10-1) P. 42 - Explain the relative physical sizes of atoms, nuclei, and electrons, showing how the relative change in volume of particles remains constant, despite the increase in their size, in terms of their mass, electric charges, and locations within the atom. (10-1) P. 42 - Explain properties of atoms and their particles (such as relative mass, electric charge) in terms of their interactions, related to P. 42, in describing the structure, organization, and function of atoms (atomic structure, atomic organization, and atomic structure).

### QUESTIONNAIRE

**Directions:** Match the atomic theory statement to its correct conclusion.

| Statement                               | Conclusion  | Correct Match |
|---|---|---------------|
| 1. Electrons are made of atoms.         | A. Atoms are made of smaller particles (protons, neutrons, electrons).  | A             |
| 2. Atoms are made of smaller particles. | B. Mass of an atom is concentrated mostly in positively charged nuclei. | B             |
| 3. Atoms are made of smaller particles. | C. Atoms are indivisible particles.                                     | C             |
| 4. Atoms are made of smaller particles. | D. Atomic number of an element is equal to the number of protons.       | D             |
| 5. Atoms are made of smaller particles. | E. Electrons have a negative charge.                                    | E             |
| 6. Atoms are made of smaller particles. | F. Atoms are electrically neutral.                                      | F             |

**Directions:** Use the periodic table to answer the following questions:

1. **Atomic number** = **P** = **20** - The number of protons equals the number of electrons.
2. **Atomic number** = **20** - The number of protons will equal the **number of electrons**.
3. **P** = **20** - **20** protons are present. Electrons are present in shells. The net charge will be **neutral**.

# Electrons In Atoms Answer Key Study Guide

**Katherine J. Denniston, Denniston  
Katherine**



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