

### Practice problems for Gravimetric analysis and gravimetric factor ratio Answers

1. Determine the gravimetric factor for the following sought analytes and weighed final products:

(a) Weighed final product:  $\text{CaC}_2\text{O}_4$

Sought analyte:  $\text{CaCO}_3$

$$\text{MW} (\text{CaC}_2\text{O}_4) = 40.08 \text{ g/mol} + 2(12.01 \text{ g/mol}) + 4(16.00 \text{ g/mol}) = 128.10 \text{ g/mol}$$

$$\text{MW} (\text{CaCO}_3) = 40.08 \text{ g/mol} + 12.01 \text{ g/mol} + 3(16.00 \text{ g/mol}) = 100.09 \text{ g/mol}$$

$$GF = \frac{1 \times \text{MW of CaCO}_3}{1 \times \text{MW of CaC}_2\text{O}_4} = \frac{100.09}{128.10} = 0.78134$$

(b) Weighed final product:  $\text{PbI}_2$

Sought analyte:  $\text{NaIO}_3$

$$\text{MW} (\text{PbI}_2) = 207.2 \text{ g/mol} + 2(126.90 \text{ g/mol}) = 461.0 \text{ g/mol}$$

$$\text{MW} (\text{NaIO}_3) = 22.99 \text{ g/mol} + 126.90 \text{ g/mol} + 3(16.00 \text{ g/mol}) = 197.89 \text{ g/mol}$$

$$GF = \frac{2 \times \text{MW of NaIO}_3}{1 \times \text{MW of PbI}_2} = \frac{2(197.89)}{461.0} = 0.8585$$

(c) Weighed final product:  $\text{B}_2\text{O}_3$

Sought analyte:  $\text{Na}_2\text{B}_4\text{O}_7$

$$\text{MW} (\text{B}_2\text{O}_3) = 2(10.81 \text{ g/mol}) + 3(16.00 \text{ g/mol}) = 69.62 \text{ g/mol}$$

$$\text{MW} (\text{Na}_2\text{B}_4\text{O}_7) = 2(22.99 \text{ g/mol}) + 4(10.81) \text{ g/mol} + 7(16.00 \text{ g/mol}) = 201.22 \text{ g/mol}$$

$$GF = \frac{1 \times \text{MW of Na}_2\text{B}_4\text{O}_7}{2 \times \text{MW of B}_2\text{O}_3} = \frac{201.22}{2(69.62)} = 1.445$$

(d) Weighed final product:  $\text{Hg}_2\text{I}_2$

Sought analyte:  $\text{HgO}$

$$\text{MW} (\text{Hg}_2\text{I}_2) = 2(200.59 \text{ g/mol}) + 2(126.90 \text{ g/mol}) = 654.98 \text{ g/mol}$$

$$\text{MW} (\text{HgO}) = 200.59 \text{ g/mol} + 16.00 \text{ g/mol} = 216.59 \text{ g/mol}$$

$$GF = \frac{2 \times \text{MW of HgO}}{1 \times \text{MW of Hg}_2\text{I}_2} = \frac{2(216.59)}{654.98} = 0.66136$$

2. The hydrogen sulphide ( $\text{H}_2\text{S}$ ) in a 50.0 g sample of crude petroleum was removed by distillation and collected in a solution of  $\text{CdCl}_2$ . The precipitated  $\text{CdS}$  was then filtered, washed, and ignited to  $\text{CdSO}_4$ . Calculate the percentage of  $\text{H}_2\text{S}$  in the sample if 0.108 g of  $\text{CdSO}_4$  was recovered.

$$\text{MW} (\text{H}_2\text{S}) = 2(1.01) \text{ g/mol} + 32.07 \text{ g/mol} = 34.09 \text{ g/mol}$$

$$\text{MW} (\text{CdSO}_4) = 112.41 \text{ g/mol} + 32.07 \text{ g/mol} + 4(16.00 \text{ g/mol}) = 208.48 \text{ g/mol}$$

$$GF = \frac{1 \times \text{MW of H}_2\text{S}}{1 \times \text{MW of CdSO}_4} = \frac{34.09}{208.48} = 0.1635$$

$$\% \text{H}_2\text{S} = [(0.108 \text{ g}) \times (0.1635)/50.0 \text{ g} \times 100] = 0.0353\%$$

# Gravimetric Analysis Problems With Answers

**Sabine Zange**



## **Gravimetric Analysis Problems With Answers:**

## The Enigmatic Realm of **Gravimetric Analysis Problems With Answers**: Unleashing the Language is Inner Magic

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web cable sizing calculation steps the process of the size calculation method consists of six steps collect data about cable

load and environmental conditions determine the minimum cable size for continuous current carrying capacity determine the minimum size of the cable based on voltage drop

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web feb 1 2023 cable size calculations calculations based on motor horse power hp and specified current input data motor voltage v l l motor horsepower cable temperature rating c ambient temperature c cable length ft and acceptable voltage drop