

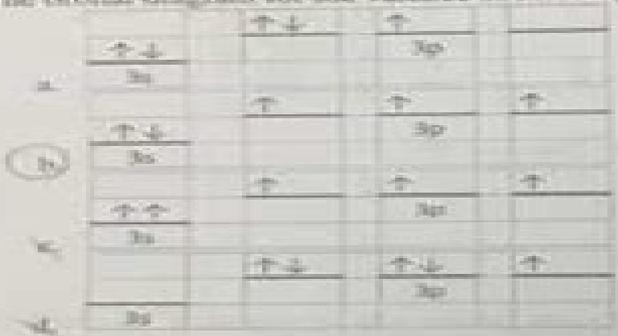
7.) Naturally-occurring copper exists as two isotopes:  $^{63}\text{Cu}$  and  $^{65}\text{Cu}$ , the masses of which are given below. Which of these isotopes, if either, is most abundant?

Isotope	Mass (amu)
$^{63}\text{Cu}$	62.929598
$^{65}\text{Cu}$	64.927730

- a. More information is needed  
 b. Neither, they are nearly-equal abundance  
 c.  $^{63}\text{Cu}$  is more abundant  
 d.  $^{65}\text{Cu}$  is more abundant

*We can find the answer given the information and they would be abundant*

8.) The orbital diagram for the valence shell of a phosphorous atom in its ground state is:



**Section II – Compounds, Reactions, Mole Concept, Solutions, and Stoichiometry**

9.) How many moles of xenon tetrafluoride correspond to 1.036 g of this compound?

- a. 2.5 mmol  
 b. 1.5 mmol  
 c. 500. mmol  
 d. 5.0 mmol  
 e. None of these

$1.036 \text{ g } \text{XeF}_4 \times \frac{1 \text{ mol}}{354.5 \text{ g}} = 0.0029 \text{ mol} \times 1000 \text{ mmol/mol} = 2.9 \text{ mmol}$

10.)

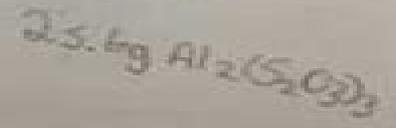
How many iodine atoms are present in 39.0 g of  $\text{Ca}(\text{IO}_3)_2$ ?

- a.  $6.002 \times 10^{23}$   
 b.  $3.011 \times 10^{23}$   
 c.  $1.20 \times 10^{24}$   
 d.  $1.20 \times 10^{22}$   
 e. None of these

$\text{Ca}(\text{IO}_3)_2 \quad 389.88 \text{ g/mol}$



$39.0 \text{ g} \times \frac{1 \text{ mol}}{389.88} \times \frac{2 \text{ mol I}}{1 \text{ mol}} = 0.200 \text{ mol I}$



$4.0 \times 10^{-1} \text{ mol} \times \frac{1 \text{ mol}}{1 \text{ mol}}$

# High School Chemistry Final Exam Answers

**R Barnett**



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