

CLASS-12 BOARDS

FORMULA SHEET

Atifa Dwaisi*

Chemistry

SOLUTIONS

Mass Percentage (w/w)% = $\frac{w_2}{w_1 + w_2} \times 100$

Volume Percentage (v/v)% = $\frac{V_2}{V_1 + V_2} \times 100$

Mass by Volume Percentage (w/v) = $\frac{w_2}{V \text{ (in ml)}} \times 100$

Mass Fraction

$X_1 = \frac{w_1}{w_1 + w_2}$ or $X_2 = \frac{w_2}{w_1 + w_2}$

Parts per million (ppm) = $\frac{w_2}{w_1 + w_2} \times 10^6$

Molarity (M) = $\frac{W_B \times 1000}{M_B \times V_{\text{soln}}}$ Unit = mol/litre

Molality (m) = $\frac{W_B \times 1000}{M_B \times W_A(g)}$ Unit = mole/kg

↳ molarity is inversely proportional to temperature

Mole Fraction (X)

$X_A = \frac{n_A}{n_A + n_B}$ $X_B = \frac{n_B}{n_A + n_B}$

$X_A + X_B = 1$ mole fraction is a unitless quantity

Henry Law

$p = K_H X_B$
↳ partial pressure of gas
↳ mole fraction of solute

RAULT'S LAW

$p_A \propto X_A$ $p_B \propto X_B$
 $p_A = p_A^0 \times X_A$ $p_B = p_B^0 \times X_B$

For Volatile Solute

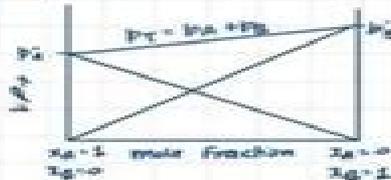
$p_r = p_A + p_B$

For Non-Volatile Solute

$p_r = p_A$
 $= p_A^0 \times X_A$

IDEAL SOLUTION

Interaction of A-B are equal to interaction of A-A and B-B



$\Delta V_{\text{mix}} = 0$
 $\Delta H_{\text{mix}} = 0$
eg n-Hexane + n-Heptane

NON-IDEAL SOLUTION

Positive Deviation

A-B interaction are weaker than A-A and B-B interaction

$\Delta V_{\text{mix}} = +ve$
 $\Delta H_{\text{mix}} = +ve$



Form minimum boiling azeotropes
eg Acetone + Ethanol

Negative Deviation

A-B interaction are stronger than A-A and B-B interaction

$\Delta V_{\text{mix}} = -ve$
 $\Delta H_{\text{mix}} = -ve$



Form maximum boiling azeotropes
eg acetone + Chloroform

Colligative Properties

depend on number of solute

Relative lowering of v.p. $\rightarrow \frac{p_A^0 - p_A}{p_A^0} = X_B$ or $i X_B$

Elevation in Boiling Point $\rightarrow \Delta T_b = K_b m$ or $i K_b m$

Depression in Freezing Point $\rightarrow \Delta T_f = K_f m$ or $i K_f m$

Osmotic Pressure $\pi = CRT$ or $\frac{n}{V} RT$ or $i \frac{n}{V} X_B RT$

Unit of K_b and $K_f = K \text{ kg mol}^{-1}$

$\Delta T_b = T_b - T_b^0$

$\Delta T_f = T_f^0 - T_f$

$\Delta T_b = K_b \frac{W_B}{M_B} \times \frac{1000}{W_A}$

$\Delta T_f = K_f \frac{W_B}{M_B} \times \frac{1000}{W_A}$

ΔT_b - Elevation in B.Pt

ΔT_f - Depression in F.Pt

T_b - B.Pt of solution

T_b^0 - B.Pt of pure solvent

T_f - F.Pt of solution

K_b - molal elevation constant

K_f - molal depression constant

Van't Hoff factor (i)

= $\frac{\text{Normal molecular mass}}{\text{observed molecular mass}}$

$i > 1$ solute undergoes dissociation

$i < 1$ solute undergoes association

$i = 1$ no association no dissociation

% dissociation = $\frac{i-1}{n-1}$ % association = $\frac{1-i}{n-1}$

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Chemistry Education Javier García-Martínez, Elena Serrano-Torregrosa, 2015-02-17 Winner of the CHOICE Outstanding Academic Title 2017 Award This comprehensive collection of top level contributions provides a thorough review of the vibrant field of chemistry education Highly experienced chemistry professors and education experts cover the latest developments in chemistry learning and teaching as well as the pivotal role of chemistry for shaping a more sustainable future Adopting a practice oriented approach the current challenges and opportunities posed by chemistry education are critically discussed highlighting the pitfalls that can occur in teaching chemistry and how to circumvent them The main topics discussed include best practices project based education blended learning and the role of technology including e learning and science visualization Hands on recommendations on how to optimally implement innovative strategies of teaching chemistry at university and high school levels make this book an essential resource for anybody interested in either teaching or learning chemistry more effectively from experience chemistry professors to secondary school teachers from educators with no formal training in didactics to frustrated chemistry students *Journal of Chemical Education*, 1925 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers

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