

Unit 5, Energetics, HL PP Pack [109 marks]

1. Which equation shows the enthalpy of formation, ΔH_f , of ethanol?

[1 mark]

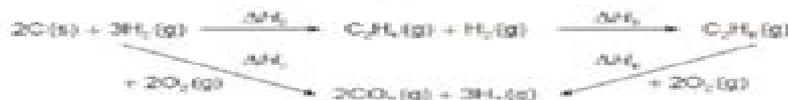
- A. $2\text{C}(\text{s}) + 3\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{C}_2\text{H}_5\text{OH}(\text{g})$
B. $4\text{C}(\text{s}) + 6\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{C}_2\text{H}_5\text{OH}(\text{g})$
C. $2\text{C}(\text{s}) + 3\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{C}_2\text{H}_5\text{OH}(\text{l})$
D. $4\text{C}(\text{s}) + 6\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{C}_2\text{H}_5\text{OH}(\text{l})$

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C

2. Which combination will give you the enthalpy change for the hydrogenation of ethene to ethane, ΔH_3 ?

[1 mark]



- A. $-\Delta H_2 + \Delta H_1 - \Delta H_4$
B. $\Delta H_2 - \Delta H_1 + \Delta H_4$
C. $\Delta H_2 + \Delta H_1 - \Delta H_4$
D. $-\Delta H_2 - \Delta H_1 + \Delta H_4$

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