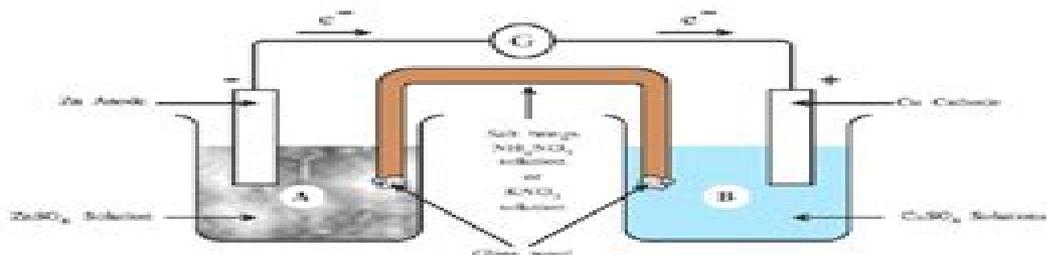


UNIT : 8 REDOX REACTIONS & ELECTROCHEMISTRY

Important Points

The device which transforms the chemical energy into electrical energy is called electrochemical cell. In the electrochemical cell, the energy producer in chemical reaction is transformed into electrical energy.



the salt bridge joins the two solutions and keeps the neutrality of electrical charge.

The cell in which the concentration of the ions associated with the reaction of the cell are 1 mole litre⁻¹ and if any gas is associated with the reaction, then its pressure equal to 1 bar and if temperature is 298 K, is called standard cell.

The relative intensity of tendency of receiving the electron is called reduction potential E_{red} or E°_{red} the relative intensity of electrodes for releasing electron is called oxidation potential

E_{red} or E°_{red}

Nernst Equation



we can write the concentrations of Cu^{2+} and Zn^{2+} ions in Daniell cell, in the given electrode potential.

$$E_{cell} = E^{\circ}_{cell} - \frac{RT}{2F} \ln \left[\frac{[Zn^{2+}]}{[Cu^{2+}]} \right]$$

Concentration Cell :

If the two electrodes of the electrochemical cell are same but the Concentration of the solution of the electrolyte are different, then it is called concentration cell. viz



Cell reaction: $Cu^{2+}(C_2) \rightleftharpoons Cu^{2+}(C_1)$ $E_{cell} = E^{\circ}_{cell} - \frac{0.059}{n} \log \left[\frac{C_1}{C_2} \right]$

Electrochemistry Practical Manual

**Dr. Erich MUELLER (of
Dresden.), Harold Johann Thomas
Ellingham**

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Laboratory Manual of Applied Electrochemistry Cornelius Sherman Grove, George Hamilton Montillon, 1939 *A Laboratory Manual of Electrochemistry* Max Erich Müller, 1931 *Laboratory Manual of Applied Electrochemistry* Oliver W. Brown, 1941 *Experimental Electrochemistry* Rudolf Holze, 2019-11-19 Showing how to apply the theoretical knowledge in practice the one and only compilation of electrochemical experiments on the market now in a new edition Maintaining its didactic approach this successful textbook provides clear and easy to follow instructions for carrying out the experiments illustrating the most important principles and applications in modern electrochemistry while pointing out the potential dangers and risks involved This second edition contains 84 experiments many of which cover electrochemical energy conversion and storage as well as electrochemical equilibrium *Handbook of Electrochemistry* Cynthia G. Zoski, 2006-12-11 Electrochemistry plays a key role in a broad range of research and applied areas including the exploration of new inorganic and organic compounds biochemical and biological systems corrosion energy applications involving fuel cells and solar cells and nanoscale investigations The Handbook of Electrochemistry serves as a source of electrochemical information providing details of experimental considerations representative calculations and illustrations of the possibilities available in electrochemical experimentation The book is divided into five parts Fundamentals Laboratory Practical Techniques Applications and Data The first section covers the fundamentals of electrochemistry which are essential for everyone working in the field presenting an overview of electrochemical conventions terminology fundamental equations and electrochemical cells experiments literature textbooks and specialized books Part 2 focuses on the different laboratory aspects of electrochemistry which is followed by a review of the various electrochemical techniques ranging from classical experiments to scanning electrochemical microscopy electrogenerated chemiluminescence and spectroelectrochemistry Applications of electrochemistry include electrode kinetic determinations unique aspects of metal deposition and electrochemistry in small places and at novel interfaces and these are detailed in Part 4 The remaining three chapters provide useful electrochemical data and information involving electrode potentials diffusion coefficients and methods used in measuring liquid junction potentials serves as a source of electrochemical information includes useful electrochemical data and information involving electrode potentials diffusion coefficients and methods used in measuring liquid junction potentials reviews electrochemical techniques incl scanning electrochemical microscopy electrogenerated chemiluminescence and spectroelectrochemistry *Handbook of Chemistry and Physics* Charles D. Hodgman, Melville F. Coolbaugh, 1917

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