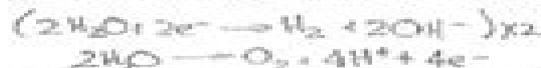


**17 • Electrochemistry****ELECTROLYSIS WORKSHEET**

Standard Reduction Potential	$E^\circ$ (volts)
$\text{Cl}_2(\text{g}) + 2\text{e}^- \rightarrow 2\text{Cl}(\text{aq})$	+1.36
$\text{O}_2(\text{g}) + 4\text{H}^+(\text{aq}) + 4\text{e}^- \rightarrow 2\text{H}_2\text{O}(\text{l})$	+1.23
$\text{Ag}^+(\text{aq}) + \text{e}^- \rightarrow \text{Ag}(\text{s})$	+0.80
$\text{I}_2(\text{s}) + 2\text{e}^- \rightarrow 2\text{I}^-(\text{aq})$	+0.535
$\text{Cu}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Cu}(\text{s})$	+0.337
$\text{SO}_4^{2-}(\text{aq}) + 4\text{H}^+(\text{aq}) + 2\text{e}^- \rightarrow \text{SO}_2(\text{g}) + 2\text{H}_2\text{O}$	+0.20
$2\text{H}^+(\text{aq}) + 2\text{e}^- \rightarrow \text{H}_2(\text{g})$ (reference electrode)	0.00
$2\text{H}_2\text{O}(\text{l}) + 2\text{e}^- \rightarrow \text{H}_2(\text{g}) + 2\text{OH}^-(\text{aq})$	-0.828
$\text{Na}^+(\text{aq}) + \text{e}^- \rightarrow \text{Na}(\text{s})$	-2.714
$\text{K}^+(\text{aq}) + \text{e}^- \rightarrow \text{K}(\text{s})$	-2.93

- All of the equations in the chart above are written as Reduction (oxidations/reductions).
- The chemicals at the upper left ( $\text{Cl}_2$  and  $\text{O}_2$ ) are the most likely to be Reduced (oxidized/reduced) and therefore the best Oxidizing agents (oxidizing agents/reducing agents).
- The chemicals at the lower right ( $\text{Na}$  and  $\text{K}$ ) are the most likely to be Oxidized (oxidized/reduced) and therefore the best Reducing agents (oxidizing agents/reducing agents).
- In an electrolytic cell, the (-) electrode is negative because it has too many (too many/too few) electrons. Chemicals that come into contact with the (-) electrode will Gain (gain/lose) electrons and be Reduced (oxidized/reduced). The (-) electrode in electrolysis is called the Cathode (cathode/anode).
- Write the change that water goes through at the (-) electrode.  $2\text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{H}_2(\text{g}) + 2\text{OH}^-(\text{aq})$
- In an electrochemical cell, the (+) electrode is positive because it has too few (too many/too few) electrons. Chemicals that come into contact with the (+) electrode will lose (gain/lose) electrons and be Oxidized (oxidized/reduced). The (+) electrode in electrolysis is called the Anode (cathode/anode).
- Write the change that water goes through at the (+) electrode.  $2\text{H}_2\text{O} \rightarrow \text{O}_2 + 4\text{H}^+ + 4\text{e}^-$
- Add these two reactions together (make certain the electrons cancel) and write the overall reaction for the electrolysis of water.  $2\text{H}_2\text{O} \rightarrow 2\text{H}_2(\text{g}) + \text{O}_2(\text{g})$
- We will perform this electrolysis using an aqueous solution of sodium sulfate. Both the  $\text{Na}^+$  and  $\text{H}_2\text{O}$  will be near the (-) electrode. Which chemical is more likely to be reduced?  $\text{H}_2\text{O}$
- Both the  $\text{SO}_4^{2-}$  and  $\text{H}_2\text{O}$  will be near the (+) electrode. Which chemical will be oxidized?  $\text{H}_2\text{O}$



# Electrolysis Experiment Lab Answer Key

**M Woodhall**



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when a yellow mongrel comes for an uninvited stay with the family travis reluctantly adopts the dog after a series of scrapes involving raccoons snakes bears wild hogs and wolves travis grows to love and respect old yeller who comes to have a profound effect on the boy s life david mullich dmullich aol com

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