

# Lab 5 Electron Configuration

CHEM101L

**Straighterline Intro to Chemistry CHEM101L Lab 5 Electron Configuration COMPLETE Worksheet with LAB PHOTOS (New Version August 2024) Scored 100%**

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**Guaranteed Passings!**

## Lab Report Format Expectations

Utilize college level grammar and professional formatting when completing this worksheet. Report all equations in a proper mathematical format, with the correct signs and symbols.

## Pre-Lab Questions

1. What is an electron configuration? Provide a response in the context of energy levels. Electron configuration is The placement of electrons within the energy levels or shells of an atom. It shows the sorts of orbitals (s, p, d, and f) and the primary quantum number (n), which represent the number of electrons occupying each energy level and sublevel. This notation aids in the prediction of the bonding arrangements and chemical activity of an element.
2. Of the following colors, which would correspond to the highest energy photons (light) and which would correspond to the lowest energy? red, orange, yellow, green, blue, indigo, violet  
The lowest energy photons are represented by red light, and the highest energy photons are represented by violet light.

## EXPERIMENT 1: OBSERVING COLORS PRODUCED FROM VARIOUS METALS

### Introduction Questions

1. In this experiment, you will conduct an experiment where atoms are heated with a flame. As the atoms gain heat from the flame, different colors of light will be produced. What is causing this light to be produced? Why are they different colors?  
The excitation and relaxing of electrons within heated atoms cause them to emit light. This phenomenon is known as flame heating. The relative energy differences between the electronic energy levels in each atom are what cause this phenomenon to emit various colors of light. Since every element has a distinct set of energy levels, distinct emission spectra are produced, which can be used to determine which elements are present. As a result, the flame test's observed colors act as a fingerprint for the components found in the sample.
2. Pick an element from the periodic table that shares a letter or two in common with the letters in your name. State the element name, its symbol, and its electron configuration below.  
sodium- Na ( $1s^2 2s^2 2p^6 3s^1$ )
3. What is the purpose of Step 6 in the experimental procedure?

**straighterline**

# Electron Configuration Lab

**JA Banks**



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