

Name : \_\_\_\_\_ Date : \_\_\_\_\_

Score : \_\_\_\_\_



## Electron Configuration



1. Determine what elements are denoted by the following configurations.

i.  $1s^2 2s^2 2p^6 3s^2 3p^4$  \_\_\_\_\_

ii.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^5$  \_\_\_\_\_

iii.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$  \_\_\_\_\_

iv.  $[\text{Kr}] 5s^2 4d^{10} 5p^3$  \_\_\_\_\_

v.  $[\text{Ne}] 3s^2 3p^1$  \_\_\_\_\_

2. Write the full electron configuration (ex.,  $1s^2 2s^1$ ) of the following elements?

i. Nitrogen \_\_\_\_\_

ii. Silicon \_\_\_\_\_

iii. Calcium \_\_\_\_\_

iv. Strontium \_\_\_\_\_

v. Bismuth \_\_\_\_\_

3. Write the abbreviated electron configuration (ex.,  $[\text{He}] 2s^1$ ) of the following element?

i. Fluorine \_\_\_\_\_

ii. Magnesium \_\_\_\_\_

iii. Selenium \_\_\_\_\_

iv. Rubidium \_\_\_\_\_

v. Lead \_\_\_\_\_

4. If each orbital can hold a maximum of two electrons, how many can each of the following hold?

i.  $2s$  \_\_\_\_\_ ii.  $5p$  \_\_\_\_\_ iii.  $4f$  \_\_\_\_\_ iv.  $3d$  \_\_\_\_\_ v.  $4d$  \_\_\_\_\_

5. What is the shape of an s orbital? \_\_\_\_\_

6. What is the shape of a p orbital? \_\_\_\_\_

# Electron Configuration Reading Worksheet

**PT Brinkman**



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