

GENERAL CHEMISTRY 2
CHAPTER TEST 2

Name: _____
Section: _____

Score: _____
Date: _____

Direction: Choose the letter of the correct answer. (*Strictly no erasures on your answer sheets.*)

- It is a special type of homogeneous mixture composed of two or more substances.
a. suspension b. **solution** c. colloids d. alloys
- Which of the following is an example of liquid in a solid solution?
a. Air b. Brass c. vinegar d. **dental amalgam**
- You are tasked to analyze a water samples containing contaminants. In what unit of concentration you are going to report the results of analysis?
a. **parts per million** b. percent c. Molarity d. molality
- What is the mole fraction of solvent in 40% isopropyl alcohol solution? (Molar mass isopropyl = 60.1 g/mol; molar mass of water = 18.0 g/mole)
a. 0.16 b. 0.67 c. **0.84** d. 3.3
- What is the concentration, in ppm, if 0.025 g KCl is dissolved in 100 grams of water?
a. 2.5×10^{-4} ppm b. 2.5 ppm c. 4.0×10^2 ppm d. **250 ppm**
- Which of the following statement is TRUE about molality?
a. Molality is temperature dependent
b. Molality is represented by moles / L
c. Molality gives a less accurate value than molarity.
d. **Molality gives a more accurate value than molarity.**
- The term homogeneous mixtures signify that
a. its properties are uniform throughout the mixture.
b. its composition is uniform throughout the mixture.
c. **Both composition and properties are uniform throughout the mixture.**
d. neither composition nor properties are uniform throughout the mixture.
- Which of the following is a quantitative description of the solution?
a. Diluted b. concentrated c. saturated d. **molality**
- Calculate the molarity of liquid HCl will be if density of the solution is 1.17 gm/cc. (MM of HCl= 36.46 g/mol; 1cc = 1 mL, 1 L= 1000 mL)
a. 31.16 M b. **32.09 M** c. 31.16 m d. 32.09 m
- What is the molarity of 20% (w/v) sulfuric acid (H_2SO_4) solution?
a. 0.51 M b. 1.02 M c. **2.04 M** d. 4.08 M
- Which of the following is used as an indicator in the titration of a weak acid and a strong base?
a. Methyl orange (3 to 4.6) c. Methyl red (5 to 6.9)
b. Bromothymol blue (6 to 7.5) d. **Phenolphthalein (8 to 9.6)**
- Arrange the procedure in based on the correct sequence of the titration process.
I. Indicator is added to the solution.
II. A specified amount of analyte is placed on the Erlenmeyer flask.
III. Titrate the sample until a distinct change in the color of the solution is observed.
IV. Titrant is placed in a burette to titrate the sample containing indicator.
a. I, II, III and IV b. II, III, IV and I c. III, IV, II and I d. **II, I, IV and III**
- The following is an indication that the end-point of titration is reached except
a. **The solution becomes dark pink in color.**
b. All the acid has been neutralized by the base.
c. A distinct change in the color of an indicator is noticeable.
d. The solution changes in color from clear to light pink solution.
- What is the concentration of the sulphuric acid solution if 50 ml of 0.5 M $Ba(OH)_2$ solution is added to 100 ml of the solution?
a. **0.25 M** b. 0.5 M c. 25 M d. 50 m
- Find the concentration of HCl, if 10 mL of 0.5 M $Ca(OH)_2$ is required to titrate 50 mL of HCl.
 $2HCl(aq) + Ca(OH)_2(aq) \rightarrow CaCl_2(aq) + 2H_2O(l)$
a. 0.1 M b. **0.2 M** c. 0.5 M d. 1.0 M
- Using the following equation:
 $2NaOH + H_2SO_4 \rightarrow 2H_2O + Na_2SO_4$
How many grams of sodium sulfate will be formed if you start with 1.25 L of a 4.0 M sodium hydroxide? (Atomic mass: Na=23.0, O=16.0, H=1.01, S=32.0.)
a. 1420 g b. 710.0 g c. **155.0 g** d. 0.7040 g

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