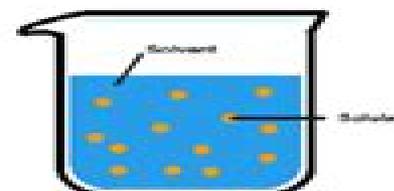


UNIT 2**SOLUTIONS**

- > A solution is a homogeneous mixture of two or more pure substances whose composition can be varied within certain limits.

BINARY SOLUTION

- > A solution consisting of two components is called a binary solution.
- > The two components are as solute and solvent.
- > The component present in excess is called the solvent.
- > The component present in lesser quantity is called solute.

**TYPES OF SOLUTIONS**

- > Classification is based on the physical state of the solute and solvent.

Types of solutions	Solute	Solvent	Examples
Solid Solutions	Solid	Solid	Copper dissolved in gold
	Liquid	Solid	Amalgam of Hg with Na
	Gas	Solid	Solution of H ₂ in Palladium
Liquid Solutions	Solid	Liquid	Glucose dissolved in water
	Liquid	Liquid	Ethanol dissolved in water
	Gas	Liquid	Oxygen dissolved in water
Gaseous Solutions	Solid	Gas	Camphor in Nitrogen gas
	Liquid	Gas	Chloroform mixed with Nitrogen gas
	Gas	Gas	Mixture of Oxygen and Nitrogen gas

METHODS FOR EXPRESSING CONCENTRATION OF SOLUTIONS

- > Concentration is the amount of solute present per unit mass of solvent or solution.

1. MASS PERCENTAGE

- > It is the mass of the component per 100g of the solution.

$$\text{Mass percentage of solute} = \frac{\text{Mass of solute}}{\text{Total mass of solution}} \times 100$$

$$\text{Mass percentage of solvent} = \frac{\text{Mass of solvent}}{\text{Total mass of solution}} \times 100$$

- > Eg: 10% glucose in water by mass means 10g of glucose is dissolved in 90g water resulting in a 100g solution.

QUESTION

What is the weight percentage of urea solution in which 10 gm of urea is dissolved in 90 gm water?

$$\text{Mass percentage of urea} = \frac{\text{Mass of urea}}{\text{Total mass of solution}} \times 100$$

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