

Released AP Multiple Choice Questions – Acids & Bases / Aqueous Equilibrium

D	33. The pH of 0.1-molar ammonia is approximately (A) 1 (B) 4 (C) 7 (D) 11 (E) 14
E	48. Which of the following ions is the strongest Lewis acid? (A) Na^+ (B) Cl^- (C) CH_3COO^- (D) Mg^{2+} (E) Al^{3+}
C	49. Each of the following can act as both a Brønsted acid and a Brønsted base EXCEPT (A) HCO_3^- (B) H_2PO_4^- (C) NH_4^+ (D) H_2O (E) HS^-
B	53. Which, if any, of the following species is in the greatest concentration in a 0.100-molar solution of H_2SO_4 in water? (A) H_2SO_4 molecules (B) H_3O^+ ions (C) HSO_4^- ions (D) SO_4^{2-} ions (E) All species are in equilibrium and therefore have the same concentrations.
E	71. Which of the following reactions does NOT proceed significantly to the right in aqueous solutions? (A) $\text{H}_3\text{O}^+ + \text{OH}^- \rightarrow 2 \text{H}_2\text{O}$ (B) $\text{HCN} + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{CN}^-$ (C) $\text{Cu}(\text{H}_2\text{O})_6^{2+} + 4 \text{NH}_3 \rightarrow \text{Cu}(\text{NH}_3)_4^{2+} + 4\text{H}_2\text{O}$ (D) $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{HSO}_4^-$ (E) $\text{H}_2\text{O} + \text{HSO}_4^- \rightarrow \text{H}_2\text{SO}_4 + \text{OH}^-$
E	75. If the acid dissociation constant, K_a , for an acid HA is 8×10^{-4} at 25°C , what percent of the acid is dissociated in a 0.50-molar solution of HA at 25°C ? (A) 0.08% (B) 0.2% (C) 1% (D) 2% (E) 4%
A	34. All of the following species can function as Brønsted-Lowry bases in solution EXCEPT (A) H_2O (B) NH_3 (C) S^{2-} (D) NH_4^+ (E) HCO_3^-

Equilibrium Ap Chemistry Multiple Choice

Thomas Griffiths



Equilibrium Ap Chemistry Multiple Choice:

Equilibrium Ap Chemistry Multiple Choice Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Equilibrium Ap Chemistry Multiple Choice**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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