

**KEY POINT**

You may have noticed by now that hydrogen does not really fit into the Periodic Table properly. Although it has 1 electron in its outer shell, it is not in Group 1 because it does not have similar properties to the other elements in Group 1: it is not a metal and does not have similar chemical properties to the alkali metals.

**EXTENSION WORK**

If you are interested you could try an internet search for 'metallic hydrogen'.

Groups

	1	2											3	4	5	6	7	0	
																			He
	H												B	C	N	O	F	Ne	
	Li	Be											Al	Si	P	S	Cl	Ar	
	Na	Mg																	
	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr	
	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe	
	Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn	
	Fr	Ra	Ac																

metal	-Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu
non-metal	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr

Figure 4.9 Metals and non-metals

**DIFFERENCES BETWEEN THE PROPERTIES OF METALS AND NON-METALS**

There are many differences between the properties of metals and non-metals.

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**L Darling-Hammond**



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