

PHET-5.2.6 Chapter 20 Assignments

Read each question of different situations involving the kinetic theory of matter, particles, and the temperature of an ideal gas. Write down the correct answer. (Remember always to give your final answer with the correct number of significant figures.)

Directions: For each assignment, you need to determine which variables are involved in the situation, as well as the relevant equations that describe the situation. Then, substitute the values provided into the equations and solve for the unknown quantities. The answers should be given with the correct number of significant figures and include the correct units. You may use the following equations and constants in your work:

1. **Pressure:** The pressure of a gas is related to the number of molecules, the volume, and the temperature. The equation that describes the relationship between these variables is given by the ideal gas law: $PV = nRT$. (Remember to use the correct units for each variable.)
2. **Temperature:** The temperature of a gas is related to the average kinetic energy of the molecules. The equation that describes the relationship between these variables is given by the equipartition theorem: $\frac{1}{2}mv^2 = \frac{3}{2}kT$. (Remember to use the correct units for each variable.)
3. **Volume:** The volume of a gas is related to the number of molecules, the pressure, and the temperature. The equation that describes the relationship between these variables is given by the ideal gas law: $PV = nRT$. (Remember to use the correct units for each variable.)

(continued)

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