

# Problem-Solving Workbook

## HOLT ChemFile

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# Holt Chemfile Problem Solving Workout Stoichiometry

**R. O'Connor, C. Mickey**



## **Holt Chemfile Problem Solving Workout Stoichiometry:**

*Holt ChemFile Lab Program*, 2005     Holt Chemistry File, 1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket     Holt Chemistry, 2003     Holt Chemistry 2 Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2003-01-01     **Chemical Problem-solving by Dimensional Analysis** Arnold B. Loebel, 1978     **Study Guide for Chemical Principles [by] Steven S. Zumdahl** Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples     **Problem Solving in Chemistry** M. Selvaratnam, Malcolm John Frazer, 1982-01-01     **Problem Exercises for General Chemistry** George Gilbert Long, Forrest C. Hentz, 1982     **Selected Solutions for Chemical Principles** Kenneth C. Brooks, Steven S. Zumdahl, 1992 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples     **Holt Chemistry** Holt Rinehart & Winston, 2003-01-24     Stoichiometry Michell J. Sienko, 1964     *A Stoichiometry Unit* David Callaghan, 2004     *Solving Problems in Chemistry* R. O'Connor, C. Mickey, 1974     **Solving Problems in Chemistry** Rod O'Connor, 1977     **Stoichiometry and Structure. Freshman Chemistry Problems and how to Solve Them** Michell Joseph Sienko, 1964     **Freshman Chemistry Problems and how to Solve Them Stoichiometry and Structure Equilibrium** Michell Joseph Sienko, 1964     **Introductory College Chemistry Students' Understanding of Stoichiometry** Adam J. Wolfer, 2000 Previous research has shown a gap between chemistry students conceptual and computational understandings of chemistry topics such as gas laws equilibrium and reactions This qualitative study examined the conceptual and computational understandings of stoichiometry of college students enrolled in a large lecture Introductory Chemistry course Factors that might influence students understandings were examined to determine their influence Possible influential factors examined included students prior coursework and their current chemistry instruction Instruction on stoichiometry was examined through classroom observations an instructor interview and review of the course resources Course exams and out of class assignments were also examined for their influence on students Student volunteers n 6 were interviewed to gauge their understanding of stoichiometry Students understanding was assessed through tasks that included a card sort solving conceptual and computational problems drawing representations of reactions and answering questions concerning their philosophy of learning chemistry Results indicated that students had an acceptable understanding of the particulate nature of matter but did not apply this knowledge to problem solving The students were most comfortable solving computational problems where they could apply algorithms learned from their instructor The students also applied algorithms in answering conceptual problems There appeared to be a connection between the students conceptual structures

of stoichiometry and their ability to solve computational problems. The lack of conceptual questions in assessment appeared to be a major contributing factor in the students' lack of conceptual understanding because the students discounted the importance of learning aspects of stoichiometry that were not included on exams. Other contributing factors included the computational focus of instruction on limiting reactant problems, textbook presentation, and student exercises.

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