

## Holt Modern Chemistry Review

### CHAPTER 5: THE PERIODIC LAW

The following pages contain the bulk (but not all) of the information for the chapter 5 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and **NOTHING** else, you should at least be able to **PASS** the test. \*\*\*\*\* Test items will be recall, examples, and/or application of this content. \*\*\*\*\*

#### OUTCOMES

- Collaborate with peer(s) to understand chemistry content (*C C*)
- Communicate chemistry content to teacher and peer(s) (*E C*)
- 5.1: Describe the organization of the modern Periodic table; distinguish between periods and groups (*T & R*)
- 5.2: Use the Periodic Table to obtain information about the properties (physical and chemical), ionization energy, and atomic sizes of the elements (*F & PK*)
- 5.2: Identify common non-metals, metalloids, and noble gases; identify the representative groups and name them (groups 1, 2, 13, 14, 15, 16, 17, 18) (*T & R*)
- 5.3: Compare periodic trends of atomic radii, ionization energy, and electronegativity (*T & R*)

#### 5.1: HISTORY OF THE PERIODIC TABLE

- **Vocabulary**
  - **periodic law** -- the law that states that the repeating chemical and physical properties of elements change periodically with the atomic numbers of the elements
  - **periodic table** -- an arrangement of the elements in order of their atomic numbers such that elements with similar properties fall in the same column
  - **lanthanides** -- a member of the rare-earth series of elements, whose atomic numbers range from 58 (cerium) to 71 (lutetium)
  - **actinides** -- any of the series of heavy radioactive elements that extends from thorium (atomic number 90) through lawrencium (atomic number 103) on the periodic table
- **Chapter Highlights**
  - The periodic law states that the physical and chemical properties of the elements are periodic functions of their atomic numbers.
  - The periodic table is an arrangement of the elements in order of their atomic numbers so that elements with similar properties fall in the same column/group/family.
  - The columns in the periodic table are referred to as groups.

#### 5.2: ELECTRON CONFIGURATION AND THE PERIODIC TABLE

- **Vocabulary**
  - **alkali metal** -- one of the elements of Group 1 of the periodic table (lithium, sodium, potassium, rubidium, cesium, and francium)
  - **alkaline-earth metal** -- one of the elements of Group 2 of the periodic table (beryllium, magnesium, calcium, strontium, barium, and radium)
  - **transition element** -- one of the metals that can use the inner shell before using the outer shell to bond
  - **main-group element** -- an element in the s-block or p-block of the periodic table
  - **halogen** -- one of the elements of Group 17 (fluorine, chlorine, bromine, iodine, and astatine); halogens combine with most metals to form salts
- **Chapter Highlights**
  - The rows in the periodic table are called periods.
  - Many chemical properties of the elements can be explained by the configurations of the elements' outermost (valence) electrons.
  - The noble gases exhibit unique chemical stability because their highest occupied levels have an octet (8) of electrons,  $?s^2?p^6$  (with the exception of helium, whose stability arises from its highest occupied level being completely filled with two electrons,  $1s^2$ ).

# Holt Chemistry Chapter 5 Quiz

**Terry C. Jones**



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