

# CHEMICAL BONDS

## Ionic Bonding

Metals and non-metals react together to form \_\_\_\_\_ compounds. They are held together by ionic bonds. Ionic bonding is a force of \_\_\_\_\_ attraction between two \_\_\_\_\_ atoms (called \_\_\_\_\_). An atom becomes charged by either donating or accepting \_\_\_\_\_ electrons from another atom. A valence electron is one from the \_\_\_\_\_ shell of the atom.

Since opposites attract, there is an attraction between the positively charged and negatively charged ions that holds them together.



If an atom \_\_\_\_\_ an electron, it becomes positively charged. This is then called a \_\_\_\_\_. \_\_\_\_\_ always donate electrons. If an atom accepts an electron, it becomes \_\_\_\_\_. This is called an \_\_\_\_\_. Non-metals always accept electrons.

Atoms do this in order to become '\_\_\_\_\_'. A stable atom is one which has obtained a \_\_\_\_\_ of electrons.

*How do we know which atoms donate and which accept electrons?*

Sodium has \_\_\_\_\_ lone electron in the outer shell. It is easiest for this atom to \_\_\_\_\_ one electron than to try and find seven more to fill the outer shell.



All \_\_\_\_\_ donate electrons and therefore become cations.

All \_\_\_\_\_ accept electrons and therefore become anions.

Chlorine has \_\_\_\_\_ electrons in its outer shell. Therefore, it is easiest for Chlorine to \_\_\_\_\_ one electron than to \_\_\_\_\_ to try and donate seven.



# Ionic And Covalent Bonds Study Guide

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## **Table of Contents Ionic And Covalent Bonds Study Guide**

1. Understanding the eBook Ionic And Covalent Bonds Study Guide
  - The Rise of Digital Reading Ionic And Covalent Bonds Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Ionic And Covalent Bonds Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ionic And Covalent Bonds Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic And Covalent Bonds Study Guide
  - Personalized Recommendations
  - Ionic And Covalent Bonds Study Guide User Reviews and Ratings
  - Ionic And Covalent Bonds Study Guide and Bestseller Lists

5. Accessing Ionic And Covalent Bonds Study Guide Free and Paid eBooks
  - Ionic And Covalent Bonds Study Guide Public Domain eBooks
  - Ionic And Covalent Bonds Study Guide eBook Subscription Services
  - Ionic And Covalent Bonds Study Guide Budget-Friendly Options
6. Navigating Ionic And Covalent Bonds Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Ionic And Covalent Bonds Study Guide Compatibility with Devices
  - Ionic And Covalent Bonds Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ionic And Covalent Bonds Study Guide
  - Highlighting and Note-Taking Ionic And Covalent Bonds Study Guide
  - Interactive Elements Ionic And Covalent Bonds Study Guide
8. Staying Engaged with Ionic And Covalent Bonds Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ionic And Covalent Bonds Study Guide
9. Balancing eBooks and Physical Books Ionic And Covalent Bonds Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ionic And Covalent Bonds Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ionic And Covalent Bonds Study Guide
  - Setting Reading Goals Ionic And Covalent Bonds Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ionic And Covalent Bonds Study Guide
  - Fact-Checking eBook Content of Ionic And Covalent Bonds Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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