

Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1ve	transfers 1ve ⁻ to F.	Li ⁺
F	7ve	accepts 1ve ⁻ from Li	F ⁻

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$ 	Li	1ve	donate 1e ⁻	Li ⁺
	Cl	7ve	accepts 1e ⁻	Cl ⁻
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$ 	Ca	2ve	donates 2e ⁻	Ca ⁺²
	O	6ve	accepts 2e ⁻	O ⁻²
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$ 	Be	2ve	donates 2e ⁻ to each F.	Be ⁺²
	F	7ve	each F accepts 1ve each	F ⁻
4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$ 	Mg	2ve	donates 2e ⁻ to S	Mg ⁺²
	S	6ve	accepts 2e ⁻	S ⁻²
5) $\text{K} + \text{F} \Rightarrow \text{KF}$ 	K	1ve	donates 1ve ⁻	K ⁺¹
	F	7ve	accepts 1ve ⁻	F ⁻

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