

Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion (-)	Polyatomic Anion Name
HClO ₃	Chloric acid	H ₃ O ⁺	ClO ₃ ⁻¹	Chlorate
H ₂ SO ₃	Sulfurous acid	2 H ₃ O ⁺	SO ₃ ⁻²	Sulfite
H ₂ SO ₄	Sulfuric acid	2 H ₃ O ⁺	SO ₄ ⁻²	Sulfate
H ₃ PO ₃	Phosphorous acid	3 H ₃ O ⁺	PO ₃ ⁻³	Phosphite
H ₃ PO ₄	Phosphoric acid	3 H ₃ O ⁺	PO ₄ ⁻³	Phosphate
HNO ₃	Nitric acid	H ₃ O ⁺	NO ₃ ⁻¹	Nitrate
HNO ₂	Nitrous acid	H ₃ O ⁺	NO ₂ ⁻¹	Nitrite
H ₂ CO ₃	Carbonic acid	2 H ₃ O ⁺	CO ₃ ⁻²	Carbonate

6. Look at the formulas of the ternary acids in Model 2.

a. How are ternary acids different from binary acids in their structure?

Ternary acids contain three elements and they also contain the element oxygen.

b. What number do you think the prefix "ter-" refers to?

"ter" refers to three

7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.

8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is "ic"

Polyatomic anion ending is "-ite" → acid name ending is "ous"

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming HClO₃?

HClO₃ has the name of chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula HCl, not HClO₃.

10. Write a rule for naming ternary acids.

*- The prefix "hydro" is not used in the name when an acid contains more than two elements and one of the elements is oxygen.
- "ate" is replaced with "ic" and "ite" is replaced with "ous".*

11. Predict the formula for chlorous acid.



12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."



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more than one method is generally used at a time The tension experiments in Chapter VIII Form a method Of investigation in which the examination of different salts shows so little variation that it appeared unnecessary to devote a separate chapter to the method The chief aim of the book is to give some account of the more important experimental work in this subject and no apology is offered for the absence Of theories of valency Chapter X Contains an account of some results besides the identification of complex compounds which have been arrived at by similar methods and which are likely to form the basis of further experiments About the Publisher Forgotten Books publishes hundreds of thousands of rare and classic books Find more at www.forgottenbooks.com This book is a reproduction of an important historical work Forgotten Books uses state of the art technology to digitally reconstruct the work preserving the original format whilst repairing imperfections present in the aged copy In rare cases an imperfection in the original such as a blemish or missing page may be replicated in our edition We do however repair the vast majority of imperfections successfully any imperfections that remain are intentionally left to preserve the state of such historical works

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