

## Unit 5 Covalent Bonding Webquest

### Activity #1- Introduction to Covalent Bonding

Open Chemical Bonding

<http://www.visionlearning.com/en/library/Chemistry/1/Chemical-Bonding/55>

Scroll down to the heading "Covalent Bonding."

1. As opposed to \_\_\_\_\_ bonding in which a complete transfer of electrons occurs, \_\_\_\_\_ bonding occurs when two or more elements \_\_\_\_\_ electrons. Covalent bonding occurs because the atoms in the compound have a similar tendency for electrons (generally to \_\_\_\_\_ electrons). This most commonly occurs when two \_\_\_\_\_ bond together. Because both of the nonmetals will want to \_\_\_\_\_ electrons, the elements involved will share electrons in an effort to \_\_\_\_\_ their valence shells.

2. How many valence electrons are in one atom of hydrogen?

3. How many valence electrons does hydrogen need to have a full first shell?

4. How does the hydrogen atom "pick up" another electron?

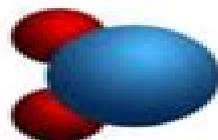
5. What compound does hydrogen form?

6. How do hydrogen atoms make a covalent bond?

7. Label the substances below as "ionic" or "covalent"



\_\_\_\_\_



\_\_\_\_\_

# Ionic And Covalent Bonding Webquest Answer Key

**Rachel Turoscy**



## **Ionic And Covalent Bonding Webquest Answer Key:**

**Ionic and Covalent Bonding** James M. Postma, Julian Roberts, J. Leland Hollenberg, 2000-01-15 **The Answer Key** Rachel Turoscy, 2018-08-09 The Answer Key A Comprehensive Explanation of Problem Solving Methods for General Chemistry Success Volume 1 is a concise and accessible textbook that covers the critical information a student needs to understand the basic mathematics used in chemistry courses The book provides easy to understand step by step instructions for solving general chemistry problems The book begins with chapters dedicated to problem solving methodology and unit conversions In subseq

**The Covalent Bond in Crystals. Ii. Partially Ionic Binding** J. C. Phillips, CHICAGO UNIV ILL., 1967 The general theory of covalent bonding developed in AD 657 432 is extended to heteropolar crystals In diatomic crystals the theory depends on two parameters an energy gap for each atom or an energy gap for the medium and a polarizability parameter  $\eta$  The latter can be determined by comparison with empirical pseudopotential form factors the value obtained for GaAs and ZnSe which brings the form factors of the latter into agreement with the ion core form factors of Animalu and Heine is  $\eta = 0.05$  or  $0.05$  This polarizability parameter can be used to estimate the effective charge  $e_c$  defined by Callen which describes the splitting of longitudinal and transverse optic modes of vibration at  $q = 0$  The result in GaAs is  $e_c = 0.14$  or  $0.04$  in rough agreement with the experimental value of  $0.20$  or  $0.03$  The experimental values for  $e_c$  may be used to estimate the gap parameters for each atom when complete data are available Author

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