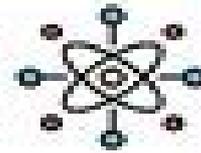


CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combination

1. **Atomic Bonding**

2. **Molecular Bonding**

3. **Network Bonding**

4. **Metallic Bonding**

CHEMICAL BONDS AND QUANTUM MECHANICS

Quantum mechanics is the study of matter and energy at the atomic and subatomic level. It explains the behavior of particles at the smallest scales, where classical physics fails. In chemistry, quantum mechanics helps us understand how atoms and molecules interact, how they form bonds, and how they react. It provides a theoretical foundation for the study of chemical bonding and the periodic table.



1. **Atomic Bonding**

2. **Molecular Bonding**

3. **Network Bonding**

4. **Metallic Bonding**

Atomic Bonding

Atomic bonding occurs when two or more atoms combine to form a molecule. The atoms are held together by the attractive forces between their positively charged nuclei and the negatively charged electrons. The resulting molecule has a lower energy state than the individual atoms, making it more stable.

Molecular Bonding

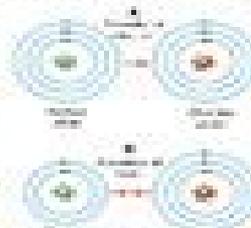
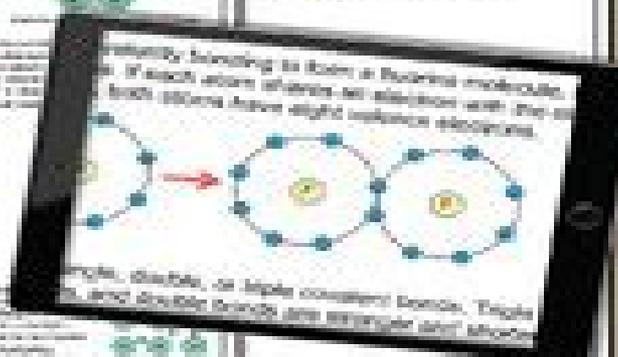
Molecular bonding occurs when two or more atoms combine to form a molecule. The atoms are held together by the attractive forces between their positively charged nuclei and the negatively charged electrons. The resulting molecule has a lower energy state than the individual atoms, making it more stable.

Network Bonding

Network bonding occurs when atoms are bonded to each other in a continuous, three-dimensional network. This type of bonding is characteristic of certain nonmetals, such as carbon, silicon, and boron. The resulting structures are very hard and have high melting points.

Metallic Bonding

Metallic bonding occurs when metal atoms are bonded to each other in a regular, repeating arrangement. The metal atoms are held together by the attractive forces between their positively charged nuclei and the delocalized electrons. This type of bonding gives metals their characteristic properties, such as malleability and conductivity.



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